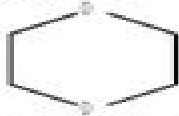


F325 Past paper Answers

1. (a) G:  $\text{CO (1)}$   
 $\text{HCOOH/H}_2\text{SO}_4 \rightarrow \text{CO} + \text{H}_2\text{O (1)}$  2
- (b) H:  $\text{C (1)}$   
 $\text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow 12\text{C} + 11\text{H}_2\text{O (1)}$  2
- (c) I:  $\text{C}_4\text{H}_8\text{O}_2 \text{ (1)}$   
 $2\text{C}_2\text{H}_4\text{O}_2 \rightarrow \text{C}_4\text{H}_8\text{O}_2 + 2\text{H}_2\text{O (1)}$   
 Structure:  
  
 (1)  
 accept any sensible structure of  $\text{C}_4\text{H}_8\text{O}_2$  3
2. (a) empirical formula  
 $\text{N} : \text{O} = 63.64/14 : 36.36/16 \text{ (1)}$   
 $= 4.56 : 2.27 = 2 : 1$ . Empirical formula =  $\text{N}_2\text{O (1)}$   
 molecular formula  
 $M_r$  of gas =  $1.833 \times 24 = 44 \text{ (1)}$  (calc 43.992)  
 with these two pieces of evidence, assume that molecular formula =  $\text{N}_2\text{O}$  3
- (b) any chemical that reacts to produce gas:  
 e.g. carbonate and  $\text{CO}_2 \text{ (1)}$   
 accept: metal more reactive than Pb and  $\text{H}_2$   
 balanced equation to match chemical added (1) 2
- (c)  $M_r(\text{Lidocaine}) = 236 \text{ (1)}$   
 moles Novocaine =  $100 \times 10^{-3}/236 = 4.24 \times 10^{-4} \text{ (1)}$   
 concentration of Novocaine =  $4.24 \times 10^{-4} \times (1000/5)$   
 $= 0.0847/0.0848 \text{ mol dm}^{-3} \text{ (1)}$  3
- (d) mass C =  $12 \times 3.74/44.0 = 12 \times 0.085 = 1.02 \text{ g (1)}$   
 mass H =  $2/18 \times 0.918 = 0.102 \text{ g (1)}$   
 mass O =  $1.394 - (1.020 + 0.102) = 0.272 \text{ g}$   
 ratio C : H : O =  $1.02/12 : 0.102/1 : 0.272/16 \text{ (1)}$   
 $= 0.0850 : 0.102 : 0.0170 / 5 : 6 : 1 / \text{C}_5\text{H}_6\text{O (1)}$   
 $\text{C}_5\text{H}_{10}\text{O}$  has relative mass of 82  
 $M_r$  is 164 =  $2 \times 82$   
 $\therefore$  molecular formula =  $\text{C}_{10}\text{H}_{12}\text{O}_2 \text{ (1)}$  5
3. (i)  $4\text{PCl}_3 + 10\text{MgO} \rightarrow \text{P}_4\text{O}_{10} + 10\text{MgCl}_2 \text{ (1)}$  1
- (ii)  $100\text{g P}_4\text{O}_{10} = \frac{100}{284} / 0.35(2) \text{ mol (1)}$   
 moles  $\text{PCl}_3$  needed =  $4 \times 0.352 = 1.408/1.4 \text{ mol (1)}$   
 mass  $\text{PCl}_3 = 1.4(08) \times 208.5 = 293.568 / 294 \text{ g/ } 291.9 \text{ g (1)}$

[7]

[13]

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**Rob Ritchie**

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