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## ACID-BASE EQUILIBRIA

### Arrhenius Acid:

Anything that produces hydrogen ions in aqueous solution

### Arrhenius Base:

Anything that produces hydroxide ions in aqueous solution

### Brønsted-Lowry Acid:

A species that donates a proton is an acid

### Brønsted-Lowry Base:

A species that accepts a proton is a base

**Amphiprotic Species:** a chemical species that can both donate and accept protons

### Lewis Acid:

Anything that accepts a pair of electrons

### Lewis Base:

Anything that donates a pair of electrons

### Water Equilibrium Constant ( $K_w$ ):

$K_w = [\text{H}_3\text{O}^+][\text{OH}^-] = 1.00 \times 10^{-14}$  (at 25°C)

### Strong Acids:

Acids that donate protons to water molecules quantitatively

### Strong Base:

A substance that generates hydroxide ions quantitatively in aqueous solution

### pH SCALES:

$\text{pH} = -\log[\text{H}_3\text{O}^+]$   $\text{pOH} = -\log[\text{OH}^-]$   $\text{p}K_a = -\log K_a$

$\text{p}K_b = -\log K_b$   $\text{pH} + \text{pOH} = 14.00$   $K_a K_b = K_w$   $\text{p}K_a + \text{p}K_b = 14.00$

### Weak Acid:

Acid that reaches equilibrium when only a small fraction of its molecules transfer protons to water

### Weak Base:

Generates hydroxide ions by accepting protons from water but reaches equilibrium when only a fraction of its molecules have done so

### Applications of Aqueous Equilibria:

#### BUFFER SOLUTION:

Contains both a weak acid and its conjugate base as major species in solution

$$\text{pH} = \text{p}K_a + \log \left( \frac{[\text{A}^-]_{\text{total}}}{[\text{HA}]_{\text{total}}} \right)$$

#### BUFFER CAPACITY:

The amount of added  $\text{H}_3\text{O}^+$  or  $\text{OH}^-$  the buffer solution can tolerate without exceeding a specified pH range

## CHEMICAL KINETICS

### REACTION MECHANISM:

The exact molecular pathway that starting materials follow on their way to becoming products

### RATE-DETERMINING STEP:

The slowest elementary step in a mechanism

### FIRST-ORDER RATE LAW:

Rate =  $k[A]$ , where A is a reactant in the overall reaction, this can be converted to

$$\ln \left( \frac{[A]_0}{[A]} \right) = kt$$

### HALF-LIFE ( $t_{1/2}$ ), 1<sup>st</sup> ORDER:

When half the original concentration has been consumed

$$[A] = 0.5[A]_0 \quad t_{1/2} = \frac{\ln 2}{k}$$

### SECOND-ORDER RATE EXPRESSION:

Rate =  $k[A]^2$ , this can be converted to

$$\frac{1}{[A]} - \frac{1}{[A]_0} = kt$$

### BI-MOLECULAR ELEMENTARY REACTION:

$A + B \rightarrow \text{products}$  Elementary rate =  $k[A][B]$

### UNI-MOLECULAR ELEMENTARY REACTION:

$C \rightarrow \text{products}$  Elementary rate =  $k[C]$

### FIRST STEP IS RATE-DETERMINING:

The predicted rate law for the overall reaction is the rate expression for that first step

### ACTIVATION ENERGY ( $E_a$ ):

Energy barrier, the minimum energy that must be supplied before the reaction can occur

**Activated Complex:** the molecular arrangement at the point of highest energy along the energy level diagram

## CHEMICAL EQUILIBRIUM

### EQUILIBRIUM CONSTANT:

$$K_{eq} = \frac{[\text{D}]_{eq}^d [\text{E}]_{eq}^e}{[\text{A}]_{eq}^a [\text{B}]_{eq}^b}$$

$K_{eq}$  applied only at equilibrium,  $K_{eq}$  is independent of initial conditions,  $K_{eq}$  is related to the stoichiometry

### Pure Liquid and Solid:

The concentrations of pure liquids or solids are always equal to their standard concentrations, therefore division by standard concentration results in a value of 1

### LARGE $K_{eq}$ :

Indicates that the reaction goes virtually to completion

### Le Chatelier's Principle:

When a change is imposed on a system at equilibrium, the system will react in the direction that reduces the amount of change

### TEMPERATURE:

The only variable that causes a change in the value of  $K_{eq}$ , an increase in temperature always shifts the equilibrium position in the endothermic direction

## THERMODYNAMICS

### CLOSED SYSTEM:

Exchanges energy but not matter with its surroundings

### ISOLATED SYSTEM:

Exchanges neither matter nor energy with the surroundings

### STATE VARIABLES:

Conditions that must be specified to establish the state of a system, pressure (P), volume (V), temperature (T), and amounts of substances (n)

### PHYSICAL CHANGE OF STATE:

Some of the state variables changes, but the chemical composition of the system stays the same

### CHEMICAL CHANGE OF STATE:

The amounts of reactants and products change

**Chemical Reaction:** some bonds break, and new bonds form

### STATE FUNCTION:

A property that depends only on the state of the system

### PATH FUNCTION:

A property that depends on how a change takes place

### Molar Heat Capacity:

The amount of heat needed to raise the temperature of 1 mol of substance by 1 Kelvin (1K)

### WORK (W):

Energy used to move an object against an opposing force,  $w = Fd$

### FIRST LAW OF THERMODYNAMICS:

$$\Delta E_{\text{sys}} = q_{\text{sys}} + w_{\text{sys}}$$

### Exothermic Process:

If the chemicals release heat, this heat gain raises the temperature of the surroundings

### Endothermic Process:

If the chemicals absorb heat, this heat loss lowers the temperature of the surroundings

$$q_{\text{surroundings}} = -C_{\text{cal}} \Delta T, \quad q = nC_s \Delta T$$

### ENTHALPY:

A thermodynamic quantity whose change equals the heat flow at constant pressure

$$\Delta H = \Delta E + \Delta(PV)$$

### STANDARD ENTHALPY OF FORMATION ( $\Delta H_f^\circ$ ):

Enthalpy change accompanying the formation of one mole of a chemical substance from pure elements in their most stable forms under standard conditions

### Hess' Law:

The enthalpy change for any overall process is equal to the sum of enthalpy changes for any set of steps that leads from the reactants to the products

### MOLAR HEAT OF SOLUTION:

Measures net energy flow that occurs as substance dissolves

### MOLAR HEAT OF VAPORIZATION:

The heat needed to vaporize one mole of a substance at its normal boiling point

### MOLAR HEAT OF FUSION:

Heat needed to melt one mole of a substance at its normal melting point

### SUBLIMATION:

A phase change in which a solid converts directly to a vapour without passing through the liquid phase

$$\Delta E_{\text{sys}} = \Delta H_{\text{sys}} - RT_{\text{sys}}$$

### SECOND LAW OF THERMODYNAMICS:

Any spontaneous process increases the disorder of the universe

### ENTROPY:

The state function that provides a quantitative measure of disorder and is symbolized S

$$\Delta S = \frac{q_r}{T}$$

### Entropy Change of the Universe:

Total entropy change

$$\Delta S_{\text{universe}} = \Delta S_{\text{system}} + \Delta S_{\text{surroundings}}$$

### REACTION ENTROPIES:

$$\Delta S_{\text{reaction}} = \sum \text{coeff}_p \Delta S^\circ - \sum \text{coeff}_r \Delta S^\circ$$



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**Yi-Tong Ma**

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