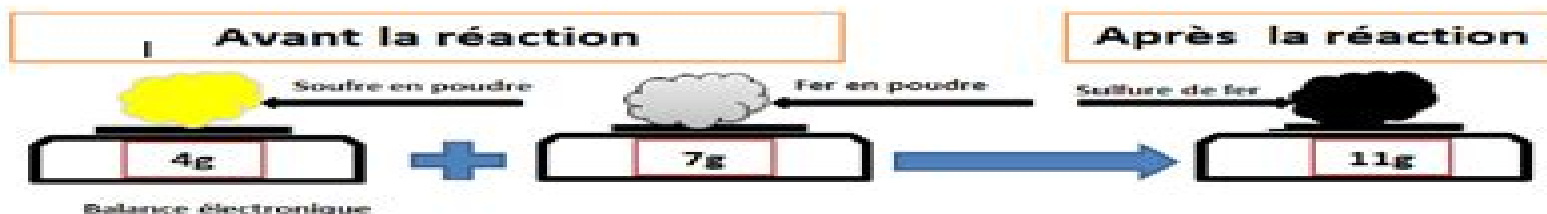


I- Loi de conservation de masse au cours d'une réaction chimique

1- Expérience

On pèse 4g de soufre et 7g de poudre de fer .Après on chauffe le mélange jusqu'à l'incandescence à l'aide du bec Bunsen:



2- Observation

Réactifs		Produits
Soufre	Fer	Sulfure de fer
$m_1 = 4\text{g}$	$m_2 = 7\text{g}$	$m = 11\text{g}$

- On observe que la masse du produit est égale à la somme des masses des réactifs
- On dit que **la masse** a été **conservée** c.-à-d. ne change pas.

3- Conclusion

Au cours d'une réaction chimique, la **masse** des réactifs qui disparaissent est **égale** à la masse des produits qui se forment : il y a **conservation de la masse**.

Exercice d'application

II- Loi de conservation des atomes au cours d'une réaction chimique

1- Exemple : Combustion du carbone dans le dioxygène

Bilan de cette réaction	Réactifs		Produits
	Carbone	Dioxygène	Dioxyde de carbone
Modèle moléculaire			
Nombre d'atome de chaque type	1 atome de carbone 2 atomes d'oxygène		1 atome de carbone 2 atomes de d'oxygène

La Classe De Rheacuteto

Clemens Wendtner



La Classe De Rheacuteto:

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