

ex: The pH of blood is 7.41. Find the pOH,  $[H^+]$ , and  $[OH^-]$ .

$$pH + pOH = 14.00$$

$$7.41 + pOH = 14.00$$

$$pOH = 6.59$$



$$pOH = -\log [OH^-]$$

$$-pOH = \log [OH^-]$$

$$10^{-pOH} = 10^{\log [OH^-]}$$

$$10^{-pOH} = [OH^-]$$

$$pH = -\log [H^+]$$

$$-pH = \log [H^+]$$

$$10^{-pH} = 10^{\log [H^+]}$$

$$10^{-pH} = [H^+]$$

"10 to the log of something equals that something."

$$10^{-7.41} = [H^+] = 3.9 \times 10^{-8} M$$

→ punch in to calculator to get ↗

$$10^{-6.59} = [OH^-] = 2.6 \times 10^{-7} M$$

### \* Calculating the pH of Strong Acid Solutions.

↳ HCl, HNO<sub>3</sub>, H<sub>2</sub>SO<sub>4</sub>, HClO<sub>4</sub>

- it is essential to know the major components (species) in the solution, and to know the minor species that can be ignored.

ex:  $[HNO_3] \rightarrow$  a very strong acid, and strong acids dissociate completely in solution.

↳ so, a 1.0M HNO<sub>3</sub> solution contains no HNO<sub>3</sub>!! Instead it contains H<sup>+</sup>, NO<sub>3</sub><sup>-</sup>, and H<sub>2</sub>O.

ex: Calculate the pH of a 0.10M HCl solution.



so 0.10M H<sup>+</sup>

↳ H<sup>+</sup>, Cl<sup>-</sup>, H<sub>2</sub>O (no "HCl")

↳ its "H<sup>+</sup>" contribution is negligible.  
( $1.0 \times 10^{-7} M H^+$ )

$$pH = -\log [H^+] = -\log (0.10) \Rightarrow pH = 1.00$$

# Gpb Chemistry Answers Solutions Key

**S Baum**



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