IGCSE - HONORS - ELECTROCHEMISTRY NOTES AND STUDY GUIDE:

10. Electrolysis of copper - refining copper

Electrodes: metallic copper

Electrolyte: Copper sulfate Solution

lons Present: Cu2+ SOy 2- OH- H+



In the Cathode Cu 24 + 2e -> Cu

In the Anode Cu -> Cu 2+ + 2e-

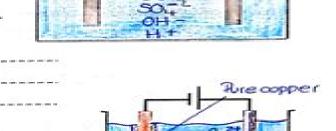
Overall Equation CU 2+ + CU -> Cu + Cu 2+

Observations:

Cathode: Dark solid deposit (pure copper)

Anode: Desintegration of anode and yellow deposit below

Uses of copper_Clectrical_wiring_capper_plating_____



What are the elements present in the stag? Aq. Au. Pt

11. Electroplating with Copper

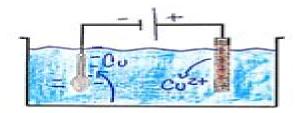
Electrodes:

Carnode: Spoon (metal more reactive than Cu)

Anode: Pure copper

Electrolyte: Copper Sulfate or Nitrate

tons Present: CU 21, SO 4 27, QH , H+



Half equations:

In the Cathode Cu2++2e -> Cu

In the Anode Cu -> Cu 2+ + 2e -

Overall Equation $C_0^{2+} + 2e^- + C_0 \rightarrow C_0 + C_0^{2+}$

Observations:

Cathoda: the color of the spoon becomes doriner

Anodo: Disintegrate Copper.

Analia Sanchez



CamScanner

Matter And Change Electrochemistry Study Guide Answer

Michael Rajnik

Matter And Change Electrochemistry Study Guide Answer:

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