

## Calculating Specific Heat Worksheet

Name: \_\_\_\_\_

Date: \_\_\_\_\_



$Q = mc\Delta T$ , where  $Q$  = heat energy,  $m$  = mass, and  $\Delta T$  = change in temp. Remember,  $\Delta T = (T_{\text{final}} - T_{\text{initial}})$ . Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$Q = (m)(c)(\Delta T)$$

$$1086.75 = (15.75)(c)(175 - 25)$$

$$1086.75 = 2362.5c$$

$$\frac{1086.75}{2362.5} = \frac{2362.5c}{2362.5}$$

$$c = 0.4672^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = (m)(c)(\Delta T)$$

$$Q = (10)(0.90)(55 - 22)$$

$$Q = 297\text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$Q = (m)(c)(\Delta T)$$

$$67500 = (1500)(c)(57 - 32)$$

$$67500 = 37500c$$

$$\frac{67500}{37500} = \frac{37500c}{37500}$$

$$c = 1.8\text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. If the specific heat of water is 4.18 J/g°C, calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mc\Delta T$$

$$Q = (100)(4.18)(37 - 4)$$

$$Q = 13827\text{ J}$$

# Mathworksheetland Answer Key

**Ensheng Dong**

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

## **Mathworksheetland Answer Key:**

## **Mathworksheetland Answer Key** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the power of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such may be the essence of the book **Mathworksheetland Answer Key**, a literary masterpiece that delves deep in to the significance of words and their impact on our lives. Written by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book is key themes, examine its writing style, and analyze its overall affect readers.

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