Chemistry 12



b) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of consumption of HCI in moles/s.

5.0×10-9 mal Mg × 2 mal HC1 = 1.0×10-8 mal HC1/s

Answer 1.0×10-8 mal HC1/s

c) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in g/s.

5.0x10-9 mol Mg x 1 mol Hz x 20 g Hz = 1.0x10-8 g Hz/s

d) If the rate of consumption of magnesium is 5.0 x 10⁻⁹ mol/s, find the rate of production of H₂ in L/s (@STP).

H2 in L/s (@STP).

5.0 x 10 - 9 mol Mg x 1 mol H2 x 22.4 L H2 = Answer 1.1 x 10 - 7 L H2/s

 e) If the rate of consumption of magnesium is 5.0 x 10-9 mol/s, find the mass of Mg consumed in 5.0 minutes.

5.0×10-9 mal Mg × 21.3 g Mg × 60 × 5.0 min = 3.6×10-5 g Mg

- When butane (C₄H₁₀) is burned in air (oxygen), the products carbon dioxide and water are formed.
 - a) Write a balanced formula equation for this reaction.

2 C4H10 +1302 -> 8CO2 + 10H20

b) If butane is consumed at an average rate of 0.116 grams/s, determine the rate of production of CO₂ in g/s.

0.116 g Cathio x 1 mol Cathio x 8 mol CO2 x 49.0 g CO2 = 0.352 g CO2/s

Answer 0.352 g CO2/s

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