

Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard potential for several cells consisting of $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes. Electrode potentials are a tool to support calculations such as those involving Gibbs free energy, Gibbs free energy of activation, and other thermodynamic quantities. Half-cell potentials are also used to determine the spontaneity of redox reactions. Other important quantities, the standard electrode potential and standard Gibbs free energy, are also used to determine the spontaneity of redox reactions. The standard electrode potential is defined as the potential of a half-cell, relative to the standard hydrogen electrode (SHE), which is assigned a potential of 0.00 V.

In this experiment, you will determine the standard electrode potential for several electrodes, which will then be used to determine the standard potential for the cell.

Objectives:

- to determine the standard electrode potential for several electrodes
- to determine the standard potential for the cell
- to determine the standard potential for the cell
- to determine the standard potential for the cell
- to determine the standard potential for the cell

Materials:

- Standard solution
- Standard solution (SHE)
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution
- Standard solution

Procedure:

- Standard solution
- Standard solution
- Standard solution
- Standard solution

Results:

- Standard solution
- Standard solution

Discussion:

- This is a good starting point.
- This is approximately 100 mV, which is the standard potential for the cell. The standard potential is approximately 0.00 V, which is the standard potential for the cell.

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E Durkheim



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