

## Thermal Energy Worksheet

1. An 80-kg patient has the flu and his temperature increases from  $37^{\circ}\text{C}$  to  $39^{\circ}\text{C}$ . How much extra heat do the chemical reactions in his cells have to generate to cause this temperature change? ( $6 \times 10^5 \text{ J}$ )

$$c = 3\,470 \text{ J/(kg}\cdot\text{K)}$$

2. Calculate

a) the amount of heat when 3.8 kg of ethylene glycol cools from  $35^{\circ}\text{C}$  to  $20^{\circ}\text{C}$ ,  $c = 2200 \text{ J/kg}\cdot\text{K}$  ( $^{\circ}\text{C} = 273^{\circ}\text{K}$ , therefore,  $1^{\circ}\text{C} = 274^{\circ}\text{K}$ ) (ans:  $1.3 \times 10^6 \text{ J}$ )

b) the temperature change when 15 kg of mercury loses 25 kJ of heat.  $c = 140 \text{ J/(kg}\cdot\text{K)}$  (ans:  $-12^{\circ}\text{K}$ )

c) the mass of copper if 7.8 kJ of heat changes its temperature from  $20^{\circ}\text{C}$  to 298 K.  $c = 390 \text{ J/(kg}\cdot\text{K)}$  (ans: 4.0 kg)

3. When 2.0 kg of a cold metal at a temperature of 248 K is immersed in 3.0 kg of water at a temperature of 313 K, the final temperature of the mixture is 309 K. What is the specific heat capacity of the metal?  $c_w = 4200 \text{ J/(kg}\cdot\text{K)}$  (ans:  $4.1 \times 10^2 \text{ J/(kg}\cdot\text{K)}$ )

Use the following table to answer the questions below:

Substance	Specific Heat Capacity $\text{J/kg}\cdot^{\circ}\text{C}$	Substance	Specific Heat Capacity $\text{J/kg}\cdot^{\circ}\text{C}$
Glass	$8.4 \times 10^2$	Water	$4.18 \times 10^3$
Iron	$4.5 \times 10^2$	Alcohol	$2.5 \times 10^3$
Brass	$3.8 \times 10^2$	Ice	$2.1 \times 10^3$
Lead	$1.3 \times 10^2$	Aluminum	$9.2 \times 10^2$
Silver	$2.4 \times 10^2$		

4. Calculate the amount of heat needed to raise the temperature of the following:

a) 8.4 kg of water by  $6.0^{\circ}\text{C}$  (ans:  $2.1 \times 10^5 \text{ J}$ )

b) 2.1 kg of alcohol by  $32^{\circ}\text{C}$  (ans:  $1.7 \times 10^5 \text{ J}$ )

5. Determine the heat lost when

a) 3.7 kg of water cools from  $31^{\circ}\text{C}$  to  $24^{\circ}\text{C}$  (ans:  $1.1 \times 10^5 \text{ J}$ )

b) a 540-g piece of silver cools from  $78^{\circ}\text{C}$  to  $14^{\circ}\text{C}$  (ans:  $8.3 \times 10^3 \text{ J}$ )

6. An electric immersion heater delivers 0.050 MJ of energy to 5.0 kg of a liquid, changing its temperature from  $32^{\circ}\text{C}$  to  $42^{\circ}\text{C}$ . Find the specific heat capacity of the liquid. (ans:  $1.0 \times 10^3 \text{ J/(kg}\cdot^{\circ}\text{C)}$ )

7. A 2.5-kg pane of glass, initially at  $41^{\circ}\text{C}$ , loses  $4.2 \times 10^4 \text{ J}$  of heat. What is the new temperature of the glass? (ans:  $21^{\circ}\text{C}$ )

# Heat And Heat Technology Directed Reading Answers

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