

16.0 g of substance in an apple. Reaction contains -2812 kJ/mole.

$$\frac{16.0 \text{ g}}{(6 \times 12.01 + 12 \times 1.008 + 6 \times 15.999) \text{ g/mole}} = 8.8110904 \times 10^{-2} \text{ mole} \times -2812 \text{ kJ/mole} \times \frac{1000 \text{ J}}{1 \text{ kJ}}$$

$$= -2.48 \times 10^5 \text{ J} \times \frac{1 \text{ Cal}}{4184 \text{ J}} = -59.6885244$$

150 g  $\text{H}_2\text{O}$  @  $25.1^\circ\text{C}$   $C_{\text{H}_2\text{O}} = 4.18 \text{ J/g}^\circ\text{C}$   $t_f = 30.1$

21.0 g  $\text{Cu}$  @  $100.4^\circ\text{C}$   $C_{\text{Cu}} = 0.385 \text{ J/g}^\circ\text{C}$   $t_f = 30.1$

$q = mc\Delta t$

$q = (150 \text{ g})(4.18 \text{ J/g}^\circ\text{C})(30.1 - 25.1)^\circ\text{C} = 3.135 \times 10^3 \text{ J} = \boxed{3.14 \times 10^3 \text{ J}}$

$q = mc\Delta t$

$q = (21.0 \text{ g})(0.385 \text{ J/g}^\circ\text{C})(100.4 - 30.1)^\circ\text{C} = \boxed{3.27 \times 10^3 \text{ J}}$

"a - b"  $(3.27 \times 10^3 - 3.135 \times 10^3) \text{ J} = 1.3455 \times 10^3 \text{ J} = 1345.5 \text{ J} = \boxed{1345 \text{ J}}$

$\Delta t$  for calorimeter is  $30.1 - 25.1 = 5^\circ\text{C} \rightarrow \frac{q}{\Delta t} = C_{\text{calor}} = \frac{1345.5 \text{ J}}{5^\circ\text{C}} = \boxed{269.1 \text{ J/}^\circ\text{C}}$

# Heat And Thermodynamics Zemansky Solution

**Michael Seilmaier**



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