

A) Describe the Kinetic Molecular Theory of Gas

Kinetic Molecular Theory

- tiny particles, atoms, or molecules moving at high velocities
- most volume of gas is empty space
- no attractive or repulsive forces between gas particles
- gas particles are constantly moving in straight paths; collisions occur with complete transfer of energy
- the average kinetic energy of gas particles is proportional to the Kelvin temp of sample

B) Identify and convert between the units of measurements used for gases

* on equation sheet

C) Use the ideal gas law to solve for P, V, T, or n of a gas

Ideal gas law: $PV = nRT$ P = pressure
V = volumeT = temperature
n = amount
(number of gas)
in moles R = ideal gas constant
= $0.0821 \frac{\text{L} \cdot \text{atm}}{\text{K} \cdot \text{mol}}$ # given
on equation
sheet

D) Use the combined gas law

$$\frac{P_1 \times V_1}{n_1 \times T_1} = \frac{P_2 \times V_2}{n_2 \times T_2}$$

E) Recognize

Inverse RelationshipP and V
T and P

$$\frac{P \times V}{n \times T}$$

Direct RelationshipV and T
n and V

$$\frac{P \times V}{n \times T}$$

F) STP

Standard Temperature + Pressure

↓
0°C
273.15 K

↓
1 atm

G) Use the molar volume at STP as a conversion factor

$$\star \frac{1 \text{ mol gas}}{22.4 \text{ L}} \quad \text{or} \quad \frac{22.4 \text{ L}}{1 \text{ mol gas}}$$

Guided Practice Problems Chapter 8 Chemistry

Clemens Wendtner



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