

6.9) It is often useful to write the equations for the reactions described:



The reaction for which the heat of reaction is requested is the formation of methane from its constituent elements.



The  $\Delta H$  for this reaction can be calculated by summing the  $\Delta H$ 's for a series of reactions which together add up to the same overall reaction (Hess's Law of Heat Summation).



These equations add up to the desired equation, so the  $\Delta H$  values are added:

$$\Delta H^\circ(\text{CH}_4) = +890. - 394 - 572 \text{ kJ/mol} = -76 \text{ kJ/mol}$$

6.10)  $\Delta H_{\text{rxn}} = -924 \text{ kJ}$

6.11) The heat of reaction is the difference between the heat of formation of the products and the heat of formation of the reactants in the balanced equation:

$$\Delta H^\circ = 12 \times (-384) + 6 \times (-242) - 2 \times (+49) = -8278 \text{ kJ.}$$

There are two moles of benzene in the equation, so the heat released for one mole is 1/2 this amount, 3139 kJ/mol.

6.12)  $\Delta H_{\text{rxn}} = -95.6 \text{ kJ}$

6.13) a. Heat lost =  $50. \text{ g} \times 0.128 \text{ J/g}^\circ\text{C} \times 68^\circ\text{C} = 260 \text{ J}$

b.  $q(\text{ice melting}) + q(\text{water warming}) = -q(\text{water cooling})$

$$5.0 \text{ g} \times 334 \text{ J/g} + 5.0 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 0^\circ\text{C}) = -100 \text{ g} \times 4.184 \text{ J/g}^\circ\text{C} (T_f - 70^\circ\text{C})$$

$$1670 \text{ J} + 20.92 T_f = -418.4 T_f + 29,288 \text{ J}$$

$$439.32 T_f = 27,618$$

$$T_f = 62.9 = 63^\circ\text{C}$$

6.14) a.  $q(\text{water}) = -q(\text{coins})$

$$c \times m \times \Delta T = -c \times m \times \Delta T$$

$$4.184 \text{ J/g}^\circ\text{C} \times 200. \text{ g} \times 5.9^\circ\text{C} = -c \times 113.4 \text{ g} \times (-74.1^\circ\text{C})$$

$$\text{Solving for } c \text{ gives: } 0.567 \text{ J/g}^\circ\text{C}$$

b. No, the specific heat capacity of pure gold is  $0.129 \text{ J/g}^\circ\text{C}$ .

6.15)  $\Delta H = 4.18 \text{ J/g}^\circ\text{C} \times 22.4 \text{ K} \times 70. \text{ g} = 6560 \text{ J}$

$$\text{For one mole of ammonium nitrate, } \Delta H = 6560 \times 60.05/20 = 26 \text{ kJ/mol.}$$

# Guided And Study Workbook Thermochemistry Answers

**Bernice G. Segal, Peter S. Shenkin**



## **Guided And Study Workbook Thermochemistry Answers:**

## **Guided And Study Workbook Thermochemistry Answers** Book Review: Unveiling the Magic of Language

In an electronic era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is truly remarkable. This extraordinary book, aptly titled "**Guided And Study Workbook Thermochemistry Answers**," compiled by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound impact on our existence. Throughout this critique, we shall delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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