

REPORT SHEET

Gravimetric Analysis of a Chloride Salt

EXPERIMENT
6

Mass of sample
Mass of filter paper + AgCl
Mass of filter paper
Mass of AgCl
Mass of Cl in original sample
(show calculations)

Trial 1	Trial 2	Trial 3
0.49	0.55	0.46
0.87	0.96	0.82
0.25	0.25	0.25
0.62	0.71	0.57
0.153356126	0.178674973	0.1409886966

$$\begin{aligned} & (0.62) \times 35.45 \text{ g Cl} \\ & (0.71) \rightarrow 143.32 \text{ g AgCl} \\ & (0.56) \rightarrow \end{aligned}$$

Percent chloride in original sample
(show calculations)

$$\frac{35.45}{(0.49)} \times 100$$

Average percent chloride (show calculations)

Standard deviation (show calculations)

Relative standard deviation expressed
as a percent (show calculations)

Do any of your results differ from the mean by more than two standard deviations?

Reported percent chloride

%

Gravimetric Analysis Of A Chloride Salt Lab Report Answers

Alice S. Cohen



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web nuclear reactions are going on all around us using correctly balanced equations is important when trying to understand nuclear reactions all equations need to be balanced to conform to two conservation laws the mass number is conserved and

the electrical charge of the nucleus is conserved