

Activity A:

Subatomic particles

Get the Gizmo ready:

- Use the arrows to create an atom with two protons, two neutrons, and two electrons.
- Turn on **Show element name**.



Question: What are the properties of protons, neutrons, and electrons?

1. **Observe:** Turn on **Show element symbol** and **Element notation**. Three numbers surround the element symbol: the **mass number** (A), electrical charge (no number is displayed if the atom is neutral), and the **atomic number** (Z).
2. **Investigate:** Watch how the numbers change as you add or remove particles.



- A. Which number is equal to the number of protons in the atom?
- B. How can you calculate the number of neutrons (N) in an atom?
- C. Which particle (proton, neutron, or electron) has a
Positive charge?
Negative charge?

neutrons

by the number of protons

electron

proton

neutron

3. **Observe:** Turn on **Isotope notation**. Isotope notation shows the number of protons and neutrons in an atom. The mass number (A) is the sum of protons and neutrons. The atomic number (Z) is the number of protons. The electrical charge is shown as a superscript to the right of the element symbol.

- A. What are the mass numbers of carbon?
- B. What are the mass numbers of oxygen?
- C. Carbon has two isotopes of oxygen.

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12, 13, 14, 15, 16, 17, 18, 19, 20, 21, 22, 23, 24, 25, 26, 27, 28, 29, 30, 31, 32, 33, 34, 35, 36, 37, 38, 39, 40, 41, 42, 43, 44, 45, 46, 47, 48, 49, 50, 51, 52, 53, 54, 55, 56, 57, 58, 59, 60, 61, 62, 63, 64, 65, 66, 67, 68, 69, 70, 71, 72, 73, 74, 75, 76, 77, 78, 79, 80, 81, 82, 83, 84, 85, 86, 87, 88, 89, 90, 91, 92, 93, 94, 95, 96, 97, 98, 99, 100

12

4. **Observe:** Turn on **Isotope notation** to answer the following questions.

- A. How many electrons are in a neutral atom of lithium?
- B. How many electrons are in an atom of hydrogen?
- C. What is the mass number of an atom with 3 protons and 3 neutrons?
- D. What is the mass number of an atom with 11 protons and 11 neutrons?
- E. How many electrons are in hydrogen?

3

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Ying-Ying Zheng

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Table of Contents Gizmos Element Builder Student Worksheet

1. Understanding the eBook Gizmos Element Builder Student Worksheet
 - The Rise of Digital Reading Gizmos Element Builder Student Worksheet
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmos Element Builder Student Worksheet
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmos Element Builder Student Worksheet
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmos Element Builder Student Worksheet
 - Personalized Recommendations
 - Gizmos Element Builder Student Worksheet User Reviews and Ratings
 - Gizmos Element Builder Student Worksheet and Bestseller Lists
5. Accessing Gizmos Element Builder Student Worksheet Free and Paid eBooks
 - Gizmos Element Builder Student Worksheet Public Domain eBooks
 - Gizmos Element Builder Student Worksheet eBook Subscription Services
 - Gizmos Element Builder Student Worksheet Budget-Friendly Options

6. Navigating Gizmos Element Builder Student Worksheet eBook Formats
 - ePub, PDF, MOBI, and More
 - Gizmos Element Builder Student Worksheet Compatibility with Devices
 - Gizmos Element Builder Student Worksheet Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmos Element Builder Student Worksheet
 - Highlighting and Note-Taking Gizmos Element Builder Student Worksheet
 - Interactive Elements Gizmos Element Builder Student Worksheet
8. Staying Engaged with Gizmos Element Builder Student Worksheet
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmos Element Builder Student Worksheet
9. Balancing eBooks and Physical Books Gizmos Element Builder Student Worksheet
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmos Element Builder Student Worksheet
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Gizmos Element Builder Student Worksheet
 - Setting Reading Goals Gizmos Element Builder Student Worksheet
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Gizmos Element Builder Student Worksheet
 - Fact-Checking eBook Content of Gizmos Element Builder Student Worksheet
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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