

4. **Think and discuss:** Why is there an attraction between the two ions in this chemical bond?

There is an attraction between these two ions in this chemical bond due to the fact that there is a positively charged atom (cation) and a negatively charged atom (anion). Opposite forces attract (+, -) and same forces repel (+, + and -, -).

Activity B: Ionic compounds	Get the Game ready: - Click Reset. Turn off Show charge. - Select Lithium (Li) and Oxygen (O).	
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Question: How are ionic compounds formed?

1. **Observe:** Look at the red lithium atom and the blue oxygen atom. Recall that most atoms are stable when their outermost ring has eight electrons. (Some atoms, such as lithium and beryllium, are stable when their outermost ring has two electrons.)

- A. How many electrons will the lithium atom give up to become stable?
- B. How many electrons does the oxygen atom need to become stable?
- C. Can a stable compound be made from these two atoms? Explain why or why not.

A stable compound cannot be made from these two atoms. Oxygen needs 2 electrons to have a stable orbit, and be happy. Though Lithium cannot give up 2 electrons without being unhappy due to the fact that it would not have a stable orbit.

2. **Experiment:** Click **Add metal** to add another lithium atom, and then transfer electrons from the lithium to the oxygen. Click **Check**.

- A. Did you make a stable compound?

- B. Turn on **Show formula**. What is the formula of this compound?

- C. Turn on **Show charge**. What is the charge of each ion? Li O

3. **Design:** Use the Game to create stable compounds from the combinations given below. After transferring electrons, arrange the atoms to demonstrate the attraction between positively charged ions and negatively charged ions. Click **Check** to check each compound.

For each compound, click the **camera** (📷) icon to take a snapshot. Right-click the image, and click **Copy Image**. Paste each image into a blank document to turn in with this worksheet. Write the ionic charges (such as Ca^{2+}) and chemical formulas below.

	Ionic charges	Chemical formula
A. Lithium and fluorine:	Li +1 F -1	LiF

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JE Gale

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'JE Gale' text.

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Table of Contents Gizmos Answers Ionic Bonds

1. Understanding the eBook Gizmos Answers Ionic Bonds
 - The Rise of Digital Reading Gizmos Answers Ionic Bonds
 - Advantages of eBooks Over Traditional Books
2. Identifying Gizmos Answers Ionic Bonds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Gizmos Answers Ionic Bonds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Gizmos Answers Ionic Bonds
 - Personalized Recommendations
 - Gizmos Answers Ionic Bonds User Reviews and Ratings
 - Gizmos Answers Ionic Bonds and Bestseller Lists
5. Accessing Gizmos Answers Ionic Bonds Free and Paid eBooks
 - Gizmos Answers Ionic Bonds Public Domain eBooks

- Gizmos Answers Ionic Bonds eBook Subscription Services
- Gizmos Answers Ionic Bonds Budget-Friendly Options
- 6. Navigating Gizmos Answers Ionic Bonds eBook Formats
 - ePub, PDF, MOBI, and More
 - Gizmos Answers Ionic Bonds Compatibility with Devices
 - Gizmos Answers Ionic Bonds Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Gizmos Answers Ionic Bonds
 - Highlighting and Note-Taking Gizmos Answers Ionic Bonds
 - Interactive Elements Gizmos Answers Ionic Bonds
- 8. Staying Engaged with Gizmos Answers Ionic Bonds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Gizmos Answers Ionic Bonds
- 9. Balancing eBooks and Physical Books Gizmos Answers Ionic Bonds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Gizmos Answers Ionic Bonds
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Gizmos Answers Ionic Bonds
 - Setting Reading Goals Gizmos Answers Ionic Bonds
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Gizmos Answers Ionic Bonds
 - Fact-Checking eBook Content of Gizmos Answers Ionic Bonds
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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