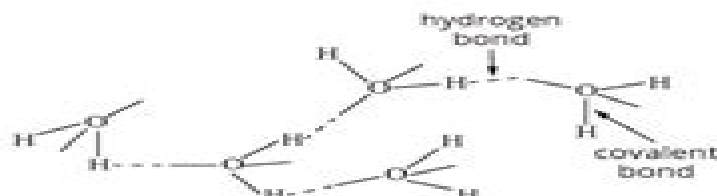


Answers to workbook worksheets

Teacher Notes

Worksheet 25 Wonderful water—structure and properties

1



- 2
- covalent
 - The O—H covalent bond is polarised because the more electronegative oxygen atom in a water molecule carries a partial negative charge and the hydrogen atoms carry a partial positive charge. The molecule is asymmetrical so it is a polar molecule.
 - see diagram for Question 1
 - Hydrogen bonding occurs between the positive partial charge on a hydrogen atom and a non-bonding pair of electrons on an oxygen atom from a different molecule.
 - examples: ammonia, hydrogen fluoride, ethanol
 - The high electronegativity of O, N and F atoms, along with their small atomic size, makes the covalent bonds they form with H atoms highly polarised. The attraction between a strong positive partial charge on the H atom and a non-bonding pair of electrons on the O, N or F atoms carrying a strong partial negative charge is relatively strong.
- 3 High latent heat of vaporisation: The water molecules can absorb a lot of energy before the hydrogen bonds between them are overcome and evaporation from liquid to gas occurs. This makes water, as sweat, a very good coolant for the human body.
- High heat capacity: Water helps to prevent overheating during exercise because the relatively strong hydrogen bonding means large amounts of energy can be absorbed by the water molecules for a relatively small increase in temperature.
- Good solvent: Nutrients for energy are delivered to the cells, and metabolic wastes are transported away and excreted, all in aqueous solution, while all of the metabolic processes for obtaining energy for exercise are aqueous reactions within cells. The ability of water to form hydrogen bonds and ion-dipole attractions to these solutes allows them to dissolve.

Worksheet 26 Specific heat capacity—calculating energy

- 1
- energy, 1°C
 - C , $\text{J g}^{-1}^{\circ}\text{C}^{-1}$, $4.18 \text{ J g}^{-1}^{\circ}\text{C}^{-1}$
 - q , $q = C \times m \times \Delta T$, specific heat capacity, mass, change in temperature
- 2
- $q = 4.18 \times 240 \times (100 - 18) = 8.23 \times 10^4 \text{ J} = 82.3 \text{ kJ}$
 - assume 3.0 L of water = 3000 g; $q = 4.18 \times 3000 \times (100 - 25) = 9.41 \times 10^5 \text{ J} = 941 \text{ kJ}$
- 3 assume 200 mL of water = 200 g; $5000 = 4.18 \times 200 \times \Delta T$, thus $\Delta T = 6.0^{\circ}\text{C}$
- 4 $1000 = 4.18 \times m \times (35 - 10)$, thus $m = 9.6 \text{ g}$

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