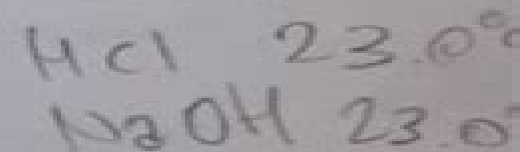


| | Reaction 1 | Reaction 2 | Reaction 3 |
|--|--------------------------------------|--|---|
| 1. Initial temperature | 22.0°C | 23.0°C | 23.0°C |
| 2. Final temperature | 24.0°C | 21.0°C | 26.0°C |
| 3. The change in temperature, ΔT | +2.0°C | -2.0°C | 3.0°C |
| 4. Heat, Q, absorbed by the solution = mass of soln. $\times \Delta T \times 1 \text{ cal/g } ^\circ\text{C}$ Note: Mass = sum of masses of solids and liquids used. | $151.6 \times 2 = 303.2 \text{ cal}$ | $160 \text{ ml} \times 4.184 \text{ J/g}^\circ\text{C} \times (-2.0^\circ\text{C}) = -1319.04 \text{ J}$ | $166.15 - 9.27 = 156.88 \text{ g}$ $156.88 \times 3.0^\circ\text{C} \times 1 = 470.64 \text{ cal}$ |
| 5. The amount of heat absorbed by the cup (neglect this quantity) | neglect | neglect | neglect |
| 6. TOTAL amount of heat, Q, absorbed | 303.2 cal | 640.0 cal | 470.64 cal |
| 7. The number of moles of NaOH used in each reaction For solid: $\text{gNaOH} \times 1/\text{MM NaOH}$ For solution: $\text{L NaOH} \times \text{Molarity}$ | 0.04 mol | 0.04 mol | 0.04 mol |
| 8. The heat of reaction per mole of NaOH, ΔH . Include sign. (Line 6/Line 7) | 7,580 cal | 16,000 cal | 11,766 cal |

(FOR EACH)

Questions

1. Write the net ionic equations for reactions 1, 2, and 3.



2. In reaction 1, ΔH_1 represents the heat of solution of NaOH(s). Look at the net ionic equations for reactions 2 and 3 and make a statement concerning the significance of ΔH_2 , that is, what has happened to produce the energy change. Also explain the meaning of ΔH_3 .

3. How does ΔH_2 compare with the sum of $\Delta H_1 + \Delta H_3$. Explain.

4. Calculate the percent difference between ΔH_2 and the SUM of $\Delta H_1 + \Delta H_3$, using ΔH_2 as the reference.

$$\frac{\text{Difference}}{\Delta H_2} \times 100\% =$$

_____ % difference

Heat Reaction Lab Answers

Didier Musso



Heat Reaction Lab Answers:

Decoding **Heat Reaction Lab Answers**: Revealing the Captivating Potential of Verbal Expression

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