

Answers

$Q = mc\Delta T$, where Q = heat energy, m = mass, and ΔT = change in temp.
Remember, $\Delta T = (T_{\text{final}} - T_{\text{initial}})$. Show all work and proper units.

1. A 15.75-g piece of iron absorbs 1086.75 joules of heat energy, and its temperature changes from 25°C to 175°C. Calculate the specific heat capacity of iron.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{1086.75}{15.75(175-25)} = 0.46 \text{ J/g}^\circ\text{C}$$

2. How many joules of heat are needed to raise the temperature of 10.0 g of aluminum from 22°C to 55°C, if the specific heat of aluminum is 0.90 J/g°C?

$$Q = mC(T_f - T_i) = 10.0\text{g} (0.90\text{J/g}^\circ\text{C})(55-22) = 297 \text{ J}$$

3. Calculate the specific heat capacity of a piece of wood if 1500.0 g of the wood absorbs 67,500 joules of heat, and its temperature changes from 32°C to 57°C.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{67500 \text{ J}}{(1500 \text{ g})(57-32)} = 1.8 \text{ J/g}^\circ\text{C}$$

4. 100.0 g of 4.0°C water is heated until its temperature is 37°C. Calculate the amount of heat energy needed to cause this rise in temperature.

$$Q = mC(T_f - T_i) = 100\text{g}(4.184\text{J/g}^\circ\text{C})(37 - 4) = 14000 \text{ J}$$

5. 25.0 g of mercury is heated from 25°C to 155°C, and absorbs 455 joules of heat in the process. Calculate the specific heat capacity of mercury.

$$C = \frac{Q}{m(T_f - T_i)} = \frac{455 \text{ J}}{(25\text{g})(155-25)} = 0.14 \text{ J/g}^\circ\text{C}$$



Heat Calculations Answer Key

J Elliott

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Table of Contents Heat Calculations Answer Key

1. Understanding the eBook Heat Calculations Answer Key
 - The Rise of Digital Reading Heat Calculations Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Heat Calculations Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Heat Calculations Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Heat Calculations Answer Key
 - Personalized Recommendations
 - Heat Calculations Answer Key User Reviews and Ratings
 - Heat Calculations Answer Key and Bestseller Lists
5. Accessing Heat Calculations Answer Key Free and Paid eBooks
 - Heat Calculations Answer Key Public Domain eBooks
 - Heat Calculations Answer Key eBook Subscription Services
 - Heat Calculations Answer Key Budget-Friendly Options
6. Navigating Heat Calculations Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Heat Calculations Answer Key Compatibility with Devices
 - Heat Calculations Answer Key Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Heat Calculations Answer Key
 - Highlighting and Note-Taking Heat Calculations Answer Key
 - Interactive Elements Heat Calculations Answer Key
8. Staying Engaged with Heat Calculations Answer Key

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Heat Calculations Answer Key
- 9. Balancing eBooks and Physical Books Heat Calculations Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Heat Calculations Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Heat Calculations Answer Key
 - Setting Reading Goals Heat Calculations Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Heat Calculations Answer Key
 - Fact-Checking eBook Content of Heat Calculations Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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