

Guided Reading (Chapter 18 Section 2)

1. What do chemical bonds do?

Chemical bonds hold the atoms of molecules together and determine their chemical and physical properties.

2. Is a covalent bond more often found in atoms that are nonmetals without variable valence?

3. Why do some molecules accept H⁺ ions in the solution? Give an example (water, alcohol, ammonia, etc.).

The lone electron pair on nitrogen in NH₃ is attracted to the H⁺ ion in H⁺.

4. What happens when a gas is cooled? Give an example. Water

5. Is water liquid at room temperature? Give an example of a substance that is.

6. What is a liquid's viscosity?

The bonds resist stretching by liquid chemical bonds, reducing viscosity.

7. Why are the water gas molecules attracted to each other?

Because there are unshared electron pairs.

8. How many electrons does the hydrogen atom have? Is hydrogen found in nature as H₂ molecules.

9. What are the high-energy bond electrons of an atom called?
valence electrons

Guided Reading Chapter 16 Section 1 Covalent

Albert A Gayle



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