

Chapter 6: Chemical Bonding Guided Notes

I. Why do Atoms Combine?

A. Chemical Formula

- The chemical formula shows:
 - The _____ in the compound.
 - The ratio of their _____



B. Chemical Bond

- Strong attractive force between _____ or ions in a molecule or compound.
- Formed by:
 - _____ electrons, (e-) by losing or gaining electrons.
 - _____ electrons (e-)

C. Stability

- Octet Rule
 - Most atoms form bonds in order to have _____ valence electrons (e-).
 - _____ outer energy level.
 - Want to be like the Noble Gases (group 18).
- Stability is the driving force behind bond formation. All atoms bond to become stable.
- Transferring e-



- Sharing e-



II. Kinds of Chemical Bonds

A. Ionic Bond

- Attraction between 2 _____ charged ions.
- Ions are _____ atoms.
- A _____ is a _____ charged ion. Forms when an atom loses an electron.
Formed by _____.
- A _____ is a _____ charged ion. Forms when an atom gains an electron.
Formed by _____.
- Formed by _____ electrons (e-) from a _____ to a _____.
- Ions form a 3-D crystal lattice.



Guided Notes Ionic Bonds Chapter 6

SJ Ball



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Table of Contents Guided Notes Ionic Bonds Chapter 6

1. Understanding the eBook Guided Notes Ionic Bonds Chapter 6
 - The Rise of Digital Reading Guided Notes Ionic Bonds Chapter 6
 - Advantages of eBooks Over Traditional Books
2. Identifying Guided Notes Ionic Bonds Chapter 6
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Guided Notes Ionic Bonds Chapter 6
 - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Notes Ionic Bonds Chapter 6
 - Personalized Recommendations
 - Guided Notes Ionic Bonds Chapter 6 User Reviews and Ratings
 - Guided Notes Ionic Bonds Chapter 6 and Bestseller Lists
5. Accessing Guided Notes Ionic Bonds Chapter 6 Free and Paid eBooks
 - Guided Notes Ionic Bonds Chapter 6 Public Domain eBooks
 - Guided Notes Ionic Bonds Chapter 6 eBook Subscription Services
 - Guided Notes Ionic Bonds Chapter 6 Budget-Friendly Options
6. Navigating Guided Notes Ionic Bonds Chapter 6 eBook Formats

- ePub, PDF, MOBI, and More
- Guided Notes Ionic Bonds Chapter 6 Compatibility with Devices
- Guided Notes Ionic Bonds Chapter 6 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Guided Notes Ionic Bonds Chapter 6
 - Highlighting and Note-Taking Guided Notes Ionic Bonds Chapter 6
 - Interactive Elements Guided Notes Ionic Bonds Chapter 6
- 8. Staying Engaged with Guided Notes Ionic Bonds Chapter 6
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Guided Notes Ionic Bonds Chapter 6
- 9. Balancing eBooks and Physical Books Guided Notes Ionic Bonds Chapter 6
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Guided Notes Ionic Bonds Chapter 6
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Guided Notes Ionic Bonds Chapter 6
 - Setting Reading Goals Guided Notes Ionic Bonds Chapter 6
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guided Notes Ionic Bonds Chapter 6
 - Fact-Checking eBook Content of Guided Notes Ionic Bonds Chapter 6
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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