

Revision - Physical Chemistry - 3b) Rates of Reaction

What 4 things increase the rate of reaction

1. Concentration

Explain using particle theory increases number of collisions so they are more likely to react

2. Temperature

Explain using particle theory increases number of collisions + the amount of energy in each collision.

3. Surface Area

Explain using particle theory increases the number of collisions so reactants are more likely to react

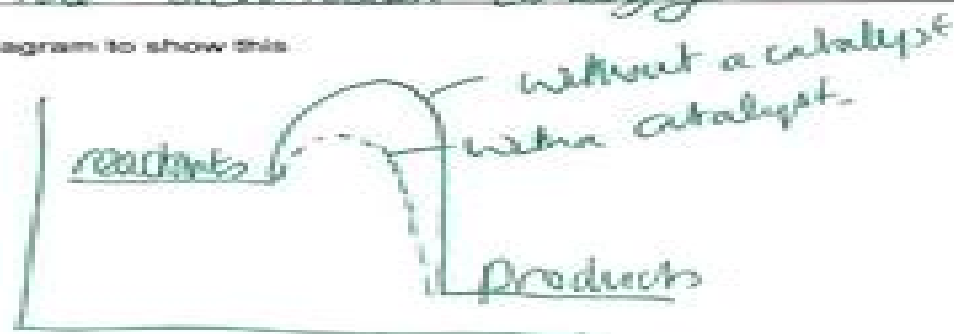
4. Pressure

Explain using particle theory decreases the space or increases particles so more frequent collisions

Define what is a catalyst speeds up a reaction without being used up

Explain how a catalyst works finds an alternative pathway to reduce the activation energy

Draw a profile diagram to show this



Guided Chemical Answers Reaction Rates And Equilibrium

Kevin Smith



Guided Chemical Answers Reaction Rates And Equilibrium:

Reviewing **Guided Chemical Answers Reaction Rates And Equilibrium**: Unlocking the Spellbinding Force of Linguistics

In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is truly astonishing. Within the pages of "**Guided Chemical Answers Reaction Rates And Equilibrium**," an enthralling opus penned by a highly acclaimed wordsmith, readers embark on an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve in to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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