

E. How many electrons are in  $\text{Mg}^{2+}$ ?

10

**Activity B:**

**Electron arrangements**

Get the Gizmo ready:

- Create a neutral hydrogen atom (1 proton, 0 neutrons, 1 electron).



**Question:** How are electrons arranged around the nucleus of an atom?

1. **Observe:** Add electrons to the atom until you have used all the available electrons. How are the electrons arranged?

They are not overlapping and are moving in circular orbits around the nucleus.

2. **Analyze:** Electrons are arranged in orbits called **energy levels**. The Gizmo shows all of the first two energy levels but only part of the third energy level.

A. How many electrons can fit in the first energy level?

2

B. How many electrons can fit in the second energy level?

8

C. How many electrons fit in the part of the third energy level shown?

8

3. **Observe:** Click **Reset** (↺). The electrons in the outermost orbit, called **valence electrons**, help to create chemical bonds. Create a lithium atom (3 protons, 4 neutrons, 3 electrons).

How many valence electrons are in a neutral lithium atom?

1

4. **Diagram:** Turn on **Show electron dot diagram**. The valence electrons of an atom are shown in an **electron dot diagram**. Each dot represents a valence electron. 🖋️ Draw the electron dot diagram for neutral lithium: (Right Click on the element, click "edit", move the dots around Li)



5. **Practice:** Turn off **Show electron dot diagram**. Use the Gizmo to create a neutral atom of each of the following elements. 🖋️ Draw the electron dot diagram for each: (Right Click on the picture, click "edit", move the dots around the element symbols) When you are finished, turn on **Show electron dot diagram** and check your answers.



6. **Extend your thinking:** Many chemical properties are determined by the number of valence electrons. Elements with the same number of valence electrons will have similar properties.

Which element has similar properties to lithium?

Na

Beryllium?

Mg

Explain:

Their valence electron count is same.

# Element Build Explore Learning Answers

**G Orfield**



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