

# 8 • Electron Configurations & Periodicity

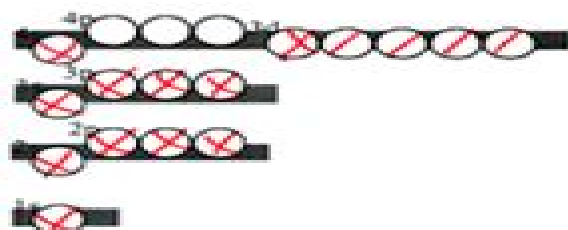
## WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
Element: Ar # of e <sup>-</sup> s: <u>18</u>	Element: Mg # of e <sup>-</sup> s: <u>12</u>	Element: N # of e <sup>-</sup> s: <u>7</u>	Element: Li # of e <sup>-</sup> s: <u>3</u>	Element: P # of e <sup>-</sup> s: <u>15</u>	Element: Cl # of e <sup>-</sup> s: <u>17</u>

Write the electron configurations of each of these in **long form** and **short form**:

- Ar  $1s^2 2s^2 2p^6 3s^2 3p^4$   
Ar  $[Ne] 3s^2 3p^4$
- Mg  $1s^2 2s^2 2p^6 3s^2$   
Mg  $[Ne] 3s^2$
- N  $1s^2 2s^2 2p^3$   
N  $[He] 2s^2 2p^3$
- Li  $1s^2 2s^1$   
Li  $[He] 2s^1$
- P  $1s^2 2s^2 2p^4 3s^2 3p^3$   
P  $[Ne] 3s^2 3p^3$
- Cl  $1s^2 2s^2 2p^6 3s^2 3p^5$   
Cl  $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



# Electron Configuration Answers

**Barron's Educational Series, Albert  
Tarendash**

A red circular graphic with a gradient, appearing as a partial circle or a thick arc, located to the right of the author's name.

## Electron Configuration Answers:

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**An Introduction to Chemistry** Michael Mosher, Paul Kelter, 2023-03-18 This textbook is written to thoroughly cover the

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## **Table of Contents Electron Configuration Answers**

1. Understanding the eBook Electron Configuration Answers
  - The Rise of Digital Reading Electron Configuration Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Answers
  - Personalized Recommendations

- Electron Configuration Answers User Reviews and Ratings
- Electron Configuration Answers and Bestseller Lists
- 5. Accessing Electron Configuration Answers Free and Paid eBooks
  - Electron Configuration Answers Public Domain eBooks
  - Electron Configuration Answers eBook Subscription Services
  - Electron Configuration Answers Budget-Friendly Options
- 6. Navigating Electron Configuration Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configuration Answers Compatibility with Devices
  - Electron Configuration Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Answers
  - Highlighting and Note-Taking Electron Configuration Answers
  - Interactive Elements Electron Configuration Answers
- 8. Staying Engaged with Electron Configuration Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Answers
- 9. Balancing eBooks and Physical Books Electron Configuration Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Answers
  - Setting Reading Goals Electron Configuration Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Answers
  - Fact-Checking eBook Content of Electron Configuration Answers

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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