

Chapter

10

ELECTROCHEMISTRY

MULTIPLE CHOICE QUESTIONS

- Reduction always takes place: (LHR 05, 12)
(a) At anode (b) At cathode
(c) At both electrodes (d) does not occur at electrodes
- The electrolyte KOH is used in cell: (GRW 06)
(a) Lead accumulator (b) Ni – Cd cell
(c) Alkaline battery (d) Silver oxide battery
- In the reaction $2\text{Fe} + 3\text{Cl}_2 \rightarrow 2\text{FeCl}_3$: (BWP 10, GRW 09, LHR 06)
(a) Fe is reduced (b) Fe is oxidized
(c) Cl_2 is oxidized (d) None of these happens
- Oxidation state of hydrogen in CaH_2 is: (GRW 07)
(a) +1 (b) -1 (c) +2 (d) zero
- Cu metal can be purified in electrolytic cell by making the impure Cu as: (GRW 07)
(a) Anode (b) Cathode (c) Anode and Cathode (d) Depends upon nature of solution
- Loss of electrons is called: (LHR 07)
(a) Oxidation (b) Reduction (c) Hydration (d) Dehydration
- Fuel cells convert chemical energy into: (GRW 07)
(a) Heat energy (b) Light energy (c) Electrical energy (d) Mechanical energy
- Electrolysis is used for: (LHR 08)
(a) Electroplating (b) Manufacture of sodium metal (c) Manufacture of Al (d) All of these
- Nelson's cell and Down's cell are example of: (FSD 09, GRW 08)
(a) Electrochemical cell (b) Galvanic cell (c) Electrolytic cell (d) None of these
- The oxidation number of Cr in $\text{K}_2\text{Cr}_2\text{O}_7$ is: (GRW 09, 10, LHR 10, 12, 13)
(a) +14 (b) +12 (c) +6 (d) +13
- The electrolyte used in fuel cell is: (LHR 09, DGK, GRW 10)
(a) Aqueous NaCl (b) Molten NaCl (c) KOH (d) H_2SO_4

Electrochemistry Multiple Multiple Choice Questions

Martha Mackin



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designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

Objective Question Bank GENERAL SCIENCE

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