

**AP® CHEMISTRY
2008 SCORING GUIDELINES**

Question 2

Answer the following questions relating to gravimetric analysis.

In the first of two experiments, a student is assigned the task of determining the number of moles of water in one mole of $\text{MgCl}_2 \cdot x\text{H}_2\text{O}$. The student collects the data shown in the following table.

Mass of empty container	22.347 g
Initial mass of sample and container	25.825 g
Mass of sample and container after first heating	23.983 g
Mass of sample and container after second heating	23.978 g
Mass of sample and container after third heating	23.977 g

- (a) Explain why the student can correctly conclude that the hydrate was heated a sufficient number of times in the experiment.

No additional mass was lost during the third heating, indicating that all the water of hydration had been driven off.

One point is earned for the correct explanation.

- (b) Use the data above to

- (i) calculate the total number of moles of water lost when the sample was heated, and

$$\text{mass of H}_2\text{O lost} = 25.825 - 23.977 = 1.848 \text{ g}$$

OR

$$25.825 - 23.906 = 1.849 \text{ g}$$

$$1.848 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.02 \text{ g H}_2\text{O}} = 0.1026 \text{ mol H}_2\text{O}$$

One point is earned for calculating the correct number of moles of water.

- (ii) determine the formula of the hydrated compound.

$$\text{mass of anhydrous MgCl}_2 = 23.977 - 22.347 = 1.630 \text{ g}$$

$$1.630 \text{ g MgCl}_2 \times \frac{1 \text{ mol MgCl}_2}{95.20 \text{ g MgCl}_2} = 0.01712 \text{ mol MgCl}_2$$

$$\frac{0.1026 \text{ mol H}_2\text{O}}{0.01712 \text{ mol MgCl}_2} = 5.993 = 6 \text{ mol H}_2\text{O per mol MgCl}_2$$

\Rightarrow Formula is $\text{MgCl}_2 \cdot 6\text{H}_2\text{O}$

One point is earned for calculating the correct number of moles of anhydrous MgCl_2 .

One point is earned for writing the correct formula (with supporting calculations).

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