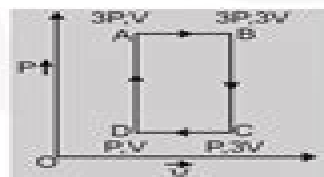


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- Pressure-temperature relationship for an ideal gas undergoing adiabatic change is ($\gamma = C_p/C_v$)
 (1) $pT^\gamma = \text{constant}$ [AIPMT 92, HARYANA PMT 01, RPET 2001]
 (2) $pT^{1-\gamma} = \text{constant}$
 (3) $pT^{1+\gamma} = \text{constant}$
 (4) $p^{1-\gamma}T^\gamma = \text{constant}$
- An inventor claims that when temperatures of source and sink are 127°C and 27°C respectively then efficiency of engine is 26%, then :- [AIPMT 2001]
 (1) It is impossible (2) It is possible but possibility is less
 (3) It's possibility is high (4) Data is insufficient
- If the system takes 100 cal. heat, and releases 80 cal to sink, if source temperature is 127°C find the sink temp. :- [AIIMS 2001]
 (1) 47°C (2) 127°C (3) 67°C (4) 107°C
- The efficiency of carnot engine is 50% and temperature of sink is 500K. If temperature of source is kept constant and its efficiency raised to 60%, then the required temperature of the sink will be:- [MANIPAL 2001, AIPMT 2002]
 (1) 100 K (2) 600 K (3) 400 K (4) 500 K
- An ideal gas heat engine operates in carnot cycle between 227°C and 127°C . It absorbs 6×10^4 cal of heat at higher temperature. Amount of heat converted to work is. [AIPMT 2005]
 (1) 2.4×10^4 cal (2) 6×10^4 cal (3) 1.2×10^4 cal (4) 4.8×10^4 cal
- An ideal monoatomic gas is taken round the cycle ABCDA as shown in following P - V diagram. The work done during the cycle is : [RPET 92, 2000, IIT 83, AIPMT 90, 92, 2007]



- (1) PV (2) 2 PV (3) 4 PV (4) Zero
- The $\left(\frac{W}{Q}\right)$ of a carnot-engine is $\frac{1}{6}$, now the temp. of sink is reduced by 62°C , then this ratio becomes twice, therefore the initial temp. of the sink and source are respectively :- [AIPMT 2007]
 (1) 33°C , 67°C (2) 37°C , 99°C
 (3) 67°C , 33°C (4) 97 K, 37 K

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