



# Missouri Trust Code And Law Manual

**KJ Lindholm-Leary**



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web q11 how do molality and molarity differ? answer both molarity and molality are concentration units. Molarity is the total number of moles of solute per litre of solution. It is defined as molarity for a given solution. The molarity of a solution is determined by several physical factors including pressure, temperature, and mass.

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web determine the molarity for each of the following solutions: 0.444 mol of COCl<sub>2</sub> in 0.654 L of solution; 98.0 g of phosphoric acid (H<sub>3</sub>PO<sub>4</sub>) in 1.00 L of solution; 0.2074 g of calcium hydroxide (Ca(OH)<sub>2</sub>) in 40.00 ml of solution.

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web calculate the mass percent, mole fraction, molality, and molarity of hexane (C<sub>6</sub>H<sub>14</sub>, d = 0.660 g/cm<sup>3</sup>) when it is mixed with heptane (C<sub>7</sub>H<sub>16</sub>, d = 0.684 g/cm<sup>3</sup>) in a 35.0 : 55.0 ml ratio. Assume that the volumes are additive.

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web for the first five problems you need to use the equation that says that the molarity of a solution is equal to the number of moles of solute divided by the number of liters of solution. In this problem, simply solve using the molarity equation to find that the concentration of the solution is 10 M.

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sulfuric acid solution containing 571.4 g of  $\text{H}_2\text{SO}_4$  per liter of solution has a density of 1.329 g  $\text{cm}^3$  calculate the molality of  $\text{H}_2\text{SO}_4$  in this solution [amazon.com chemistry and our universe how it all works](#)

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web problem 6.1.6.1.6 calculate the molarity of each of the following solutions a 0.195 g of cholesterol  $\text{C}_{27}\text{H}_{46}\text{O}$  in 0.100 l of serum the average concentration of cholesterol in human serum b 4.25 g of  $\text{NH}_3$  in 0.500 l of solution the concentration of  $\text{NH}_3$  in household ammonia

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web a similar unit of concentration is molality  $m$  which is defined as the number of moles of solute per kilogram of solvent not per liter of solution  $\text{molality} = \frac{\text{moles solute}}{\text{kilograms solvent}}$  mathematical manipulation of molality is the

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web 1 molarity  $m$   $\text{NaCl}$  mass molar mass  $\times 0.4500 \text{ l} = 0.825 \text{ g}$   $\frac{0.825 \text{ g}}{58.44 \text{ g/mol}} = 0.0141 \text{ mol}$   $\frac{0.0141 \text{ mol}}{0.4500 \text{ kg}} = 0.0313 \text{ m}$  2 molality  $m$   $\text{NaCl}$   $\frac{0.825 \text{ g}}{58.44 \text{ g/mol}} = 0.0141 \text{ mol}$   $\frac{0.0141 \text{ mol}}{0.4500 \text{ kg}} = 0.0313 \text{ m}$  3 mole fraction  $x_{\text{NaCl}} = \frac{0.0141 \text{ mol}}{0.0141 \text{ mol} + 0.4500 \text{ kg} \times \frac{1000 \text{ g}}{18.015 \text{ g/mol}}} = 0.0031$

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web molality differs from molarity only in the denominator while molarity is based on the liters of solution molality is based on the kilograms of solvent concentrations expressed in molality are used when studying properties of solutions related to vapor pressure and temperature changes

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web what mass of  $\text{CoCl}_2$  in grams is needed for the solution express the answer using 3 significant figures

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web the hardness of water hardness count is usually expressed in parts per million by mass of  $\text{CaCO}_3$  which is equivalent to milligrams of  $\text{CaCO}_3$  per liter of water what is the molar concentration of  $\text{Ca}^{2+}$  ions in a water sample with a hardness count of 175 mg  $\text{CaCO}_3/\text{l}$

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web answer the mass percentage of an element of a solution is defined as the mass of the solute in grammes present in 100 g of the solution it is written like this  $\text{begin array l} \frac{\text{mass of a component}}{\text{mass of component in solution total mass of solution}} \times 100 \text{ end array}$

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