

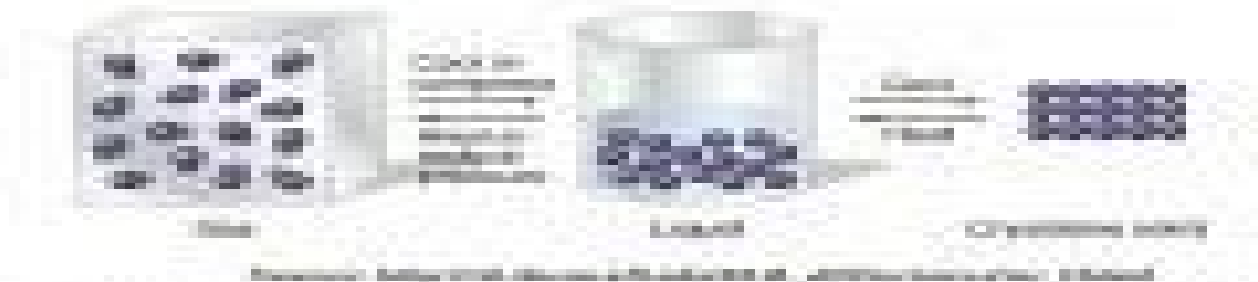
# PHYSICS: Kinetic Molecular Theory

## Learning Objectives

- 1) Identify/Describe the difference between particles/molecules in different phases (solid, liquid, gas) and how they move.
- 2) Explain the relationship between the pressure and the average kinetic energy.
- 3) Explain how temperature affects the pressure in a system.
- 4) Explain how pressure affects the volume of a gas at a constant temperature.

## Module 1: Representation of Molecules in Different Phases

The following figure shows the representation of particles in three different phases of matter:



Remember, temperature is a measure of the average kinetic energy of the particles. Pressure is a measure of the force exerted by the particles on the walls of the container.

## Key Concepts

- 1. Matter can be in three different states: solid, liquid, and gaseous. The particles in each state are arranged differently. In the solid state, particles are packed closely together. In the liquid state, particles are in a disordered, close-packed arrangement. In the gaseous state, particles are in a disordered, widely spaced arrangement.

Table 1: Characteristics of the Three States of Matter

	SOLID	LIQUID	GAS
<b>How are the particles arranged?</b>	Particles are packed closely together in a regular, ordered arrangement. They are held together by strong attractive forces.	Particles are in a disordered, close-packed arrangement. They are held together by moderate attractive forces.	Particles are in a disordered, widely spaced arrangement. They are held together by very weak attractive forces.
<b>How do they move?</b>	Particles vibrate about their fixed positions. They cannot move from one place to another.	Particles can move from one place to another. They are in constant motion.	Particles move rapidly in all directions. They are in constant motion.

# Lw Pogil Kinetic Molecular Theory Answers

**Allan Douglas Risteen**



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*Science Citation Index*, 1993 Vols for 1964 have guides and journal lists *Applications of the Kinetic Theory to Gases, Vapors, Pure Liquids, and the Theory of Solutions* William Pingry Boynton, 1904 *Molecules and the Molecular Theory of Matter* Allan Douglas Risteen, 1896 **The Molecular Theory of Solutions** I. Prigogine, 2003-01 The Molecular Theory of Solutions Ilya Prigogine, 1957 **The Potential Distribution Theorem and Models of Molecular Solutions** Tom L. Beck, Michael E. Paulaitis, Lawrence R. Pratt, 2006-08-31 An understanding of statistical thermodynamic molecular theory is fundamental to the appreciation of molecular solutions This complex subject has been simplified by the authors with down to earth presentations of molecular theory Using the potential distribution theorem PDT as the basis the text provides a discussion of practical theories in conjunction with simulation results The authors discuss the field in a concise and simple manner illustrating the text with useful models of solution thermodynamics and numerous exercises Modern quasi chemical theories that permit statistical thermodynamic properties to be studied on the basis of electronic structure calculations are given extended development as is the testing of those theoretical results with ab initio molecular dynamics simulations The book is intended for students taking up research problems of molecular science in chemistry chemical engineering biochemistry pharmaceutical chemistry nanotechnology and biotechnology **Molecular Theory of Water and Aqueous Solutions** Arieh Ben-Naim, 2009 Molecular Theory of Solutions Arieh Ben-Naim, 2006-07-27 This book presents new and updated developments in the molecular theory of mixtures and solutions It is based on the theory of Kirkwood and Buff which was published more than fifty years ago This theory has been dormant for almost two decades It has recently become a very powerful and general tool to analyze study and understand any type of mixtures from the molecular or the microscopic point of view The traditional approach to mixture has been for many years based on the study of excess thermodynamic quantities This provides a kind of global information on the system The new approach provides information on the local properties of the same system Thus the new approach supplements and enriches our information on mixtures and solutions

**Foundations of the Molecular Theory** John Dalton, Amedeo Avogadro, Joseph Louis Gay-Lussac, 1911 **Molecular Theory of Water and Aqueous Solutions** Arieh Ben-Naim, 2011 The aim of this book is to explain the unusual properties of both pure liquid water and simple aqueous solutions in terms of the properties of single molecules and interactions among small numbers of water molecules It is mostly the result of the author's own research spanning over 40 years in the field of aqueous solutions Jacket Avogadro and Dalton Andrew Norman Meldrum, 1906 **Kinetic Molecular Theory** Charles Proteus Steinmetz, American Electrochemical Society. Meeting, 1908\* **Foundations of the molecular theory, comprising papers and extracts by John Dalton, Joseph-Louis Gay-Lussac, and Amedeo Avogadro (1808-1811)** John Dalton, 1911 **Foundations of the Molecular Theory** John Dalton, Amedeo Avogadro, 2017-07-25 From the Preface THE papers here reprinted in chronological order serve to exhibit the historical development of the idea of a connection existing

between the number of particles in different gases and the volume they occupy It will be seen that Dalton from the first entertains the notion that equal volumes of different gases may contain the same number of ultimate particles at equal temperature and pressure but that he is legitimately forced to reject this assumption conceiving no distinction between the atom and the molecule of an element Gay Lussac s important experimental work on the combining volumes of gases then shows the necessity of a simple relation between the ultimate particles of gases and their volumes although he does not point this out in his paper Dalton however perceives the necessity and characteristically concludes by doubting the accuracy of Gay Lussac s experiments Avogadro finally accepts both Dalton s theory and Gay Lussac s data and teaches how to reconcile them by distinguishing between the atom and the molecule of an elementary gas It has not been thought necessary to reprint the letter of Ampere to Berthollet Annales de Chimie 90 43 86 1814 since that paper contains no advance on the views of Avogadro published three years earlier its author simply drawing the same conclusions from the same premises The English version of the French originals will probably be found more faithful than elegant especially so in the case of Avogadro s paper where the French is always clumsy and occasionally obscure

Molecular Theory of Gases and Liquids Joseph Oakland Hirschfelder,2003-01

Foundations of Molecular Theory, Comprising Papers and Extracts John Dalton,Joseph Louis Gay-Lussac,Amedeo Avogadro,1961

**Foundations of the Molecular Theory**,1893

**The Potential Distribution Theorem and Models of Molecular Solutions** Thomas L. Beck,Michael E. Paulaitis,Lawrence Riley Pratt,2006 Modern quasi chemical theories that permit statistical thermodynamic properties to be studied on the basis of electronic structure calculations are given extended development as is the testing of those theoretical results with ab initio molecular dynamics simulations The book is intended for students undertaking research problems of molecular science in chemistry chemical engineering biochemistry pharmaceutical chemistry nanotechnology and biotechnology

BOOK JACKET

**MOLECULES AND THE MOLECULAR THEORY OF MATTER** A. D. RISTEEN,2018

**Foundations of the Molecular Theory Comprising Papers and Extracts** John Dalton,Joseph Louis Gay-Lussac,Amedeo Avogadro,1923

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