

PART A – CHEMISTRY

1. An unknown alcohol is treated with the "Lucas reagent" to determine whether the alcohol is primary, secondary or tertiary. Which alcohol reacts fastest and by what mechanism:
 (1) tertiary alcohol by S_N1 (2) secondary alcohol by S_N2
 (3) tertiary alcohol by S_N2 (4) secondary alcohol by S_N1

Sol. (1)
 Reaction proceeds through carbocation formation as 3° carbocation is highly stable, hence reaction proceeds through S_N1 with 3° alcohol.

2. The first ionization potential of Na is 5.1 eV. The value of electron gain enthalpy of Na^+ will be:
 (1) - 5.1 eV (3) -10.2 eV
 (2) + 2.55 eV (4) - 2.55 eV

Sol. (1)
 $Na \xrightleftharpoons[\text{IE} = 5.1 \text{ eV}]{\text{EGE} = -5.1 \text{ eV}} Na^+ + e^-$, here the backward reaction releases same amount of energy and known as Electron gain enthalpy.

3. Stability of the species Li_2 , Li_2^+ and Li_2^- increases in the order of:
 (1) $Li_2^- < Li_2^+ < Li_2$ (3) $Li_2 < Li_2^- < Li_2^+$
 (2) $Li_2^+ < Li_2 < Li_2^-$ (4) $Li_2 < Li_2^+ < Li_2^-$

Sol. (1)
 $Li_2 (6) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^2$
 B.O. = $\frac{4-2}{2} = 1$
 $Li_2^+ (5) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^1$
 B. O. = $\frac{3-2}{2} = 0.5$
 $Li_2^- (7) = \sigma 1s^2 \sigma^* 1s^2 \sigma 2s^2 \sigma^* 2s^1$
 B.O. = $\frac{4-3}{2} = 0.5$
 Li_2^+ is more stable than Li_2^- because Li_2^+ has more numbers of antibonding electrons.

4. The molarity of a solution obtained by mixing 750 mL of 0.5 (M) HCl with 250 mL of 2(M)HCl will be:
 (1) 1.00 M (2) 1.75 M
 (3) 0.975 M (4) 0.875 M

Sol. (4)
 $M_1V_1 + M_2V_2 = MV$
 $M = \frac{M_1V_1 + M_2V_2}{V} = \frac{0.5 \times 750 + 2 \times 250}{1000}$
 $M = 0.875$

5. Which of the following is the wrong statement?
 (1) O_3 molecule is bent (2) Ozone is violet-black in solid state
 (3) Ozone is diamagnetic gas (4) $ONCl$ and $ONCl^-$ are not isoelectronic

Sol. (All the options are correct statements)

(1) Correct, as  is bent.

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Rachel S Tattersall



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your dormant account will be reactivated immediately

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