



Social Studies Chapter 3: Regions Around You Study Guide

A. Lesson One: "Understanding Regions"

A. region	B. ethnic group	C. technology	D. communication
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Directions: Match the words above to their definition below.

- _____ 1. made up of people from the same country or people with a shared way of life
- _____ 2. the way people use knowledge and tools to make or do something
- _____ 3. an area with at least one feature that makes it different from other areas
- _____ 4. the way people send and receive information

Directions: Choose the best answer and write it on the blank.

- _____ 5. Why did we divide our world into regions?
- a. We know it'd be easier to make maps if we had regions.
 - b. There are many features in our world and it would be difficult to describe them all, so we divided the world into regions.
 - c. We divided the world into regions because it's easier to travel that way.
 - d. Because we needed leaders, we divided our world into regions.
- _____ 6. Each place in the United States is part of _____.
- a. the solar system
 - b. a book
 - c. many kinds of regions
 - d. a club

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 $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) + \text{C}_2\text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g})$ $\Delta h = 94.5 \text{ kJ}$
 $\text{H}_2\text{O}(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 71.2 \text{ kJ}$
 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta h = 283 \text{ kJ}$

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web chemistry 120 hess s law worksheet calculate Δh for the reaction $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$ from the following data
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 2 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1411 \text{ kJ}$ mole
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{l})$ $\Delta h = 1560 \text{ kJ}$ mole
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\Delta h = 285 \text{ kJ}$ mole

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 hess s law example 12.1 determine Δh for the target reaction
 $\text{C}_2\text{H}_4(\text{g}) + \text{O}_2(\text{g}) \rightarrow \text{C}_2\text{H}_4\text{O}(\text{g})$
 given the following information

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 $\text{Mg}(\text{s}) + \frac{1}{2} \text{O}_2(\text{g}) \rightarrow \text{MgO}(\text{s})$ we used

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 cohen 2016 the equations for the reactions in the experiment done are as follows
 $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq})$

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web thus for reaction 9 2 in which 1 mole of H_2SO_4 and 1 mole of $NaOH$ combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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solution eventually allowing an individual to prove or disprove Hess's Law. Introduction: Thermodynamics is the study of energy exchanged between the system and the surroundings.

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Web 3 12 Hess's Law: In this experiment, you will measure the amount of heat released in these three related exothermic reactions:

1. $\text{NaOH (s)} \rightarrow \text{Na}^+ \text{(aq)} + \text{OH}^- \text{(aq)}$
2. $\text{NaOH (s)} + \text{H}^+ \text{(aq)} \rightarrow \text{Na}^+ \text{(aq)} + \text{H}_2\text{O (l)}$
- 3.