

TOPIC 8 : HEAT & TEMPERATURE

Specific Heat Capacity :-

$$\Delta E = mc\Delta\theta$$

where; c = specific heat capacity

m = mass

$\Delta\theta$ = change in temperature

Specific Latent Heat :-

$$\Delta E = L\Delta m$$

L = specific latent heat

Kinetic Energy of molecule in a gaseous sample

$$\frac{1}{2}m\langle c^2 \rangle = \frac{3}{2}kT$$

where; k = Boltzmann constant
 $1.38 \times 10^{-23} \text{ J K}^{-1}$

T = temperature in Kelvin

GIAS LAWS

• $V \propto T$

$$\frac{V_1}{T_1} = \frac{V_2}{T_2}$$

• $p \propto T$

• $p \propto \frac{1}{V}$ $p_1V_1 = p_2V_2$

where, p = pressure

T = temp.

V = vol.

Gras laws ultimate equation

$$pV = NkT$$

where; N = number of molecules in gas

k = Boltzmann const.

T = temp. in Kelvin

Equation of state for ideal gas

$$pV = nRT$$

$$pV = \frac{1}{3}mN\langle c^2 \rangle$$

where; n = number of moles of gas

R = universal gas constant = $8.31 \text{ J kg}^{-1} \text{ mol}^{-1}$

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