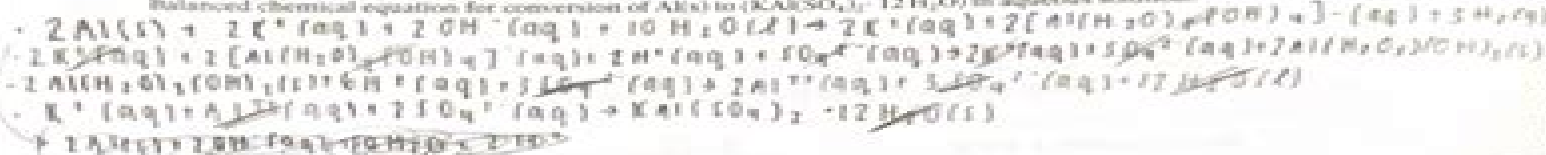


## Sample Calculations

Balanced chemical equation for conversion of Al(s) to  $(\text{KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O})$  in aqueous solution:

## Theoretical Yield:

• Moles alum • MW alum (g/mol) = grams of alum

$$0.83 \text{ g Al} \left( \frac{1 \text{ mol Al}}{26.982 \text{ g Al}} \right) \left( \frac{2 \text{ mol KAl}(\text{SO}_4)_2 \cdot 12 \text{H}_2\text{O}}{2 \text{ mol Al}} \right) \left( \frac{474.39 \text{ g alum}}{1 \text{ mol alum}} \right) = 14.59 \text{ g}$$

## Percent Yield:

$$\frac{\text{actual}}{\text{theo}} \times 100$$

$$\frac{11.56}{14.59} \times 100 = 79.23\%$$



# It In The Cards Chemistry Lab Answers

**John O'Donoghue**



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