

UNIT (6) ACIDS AND BASES

6.1 | Arrhenius Definition of Acids and Bases

Definitions for acids and bases were proposed by the Swedish chemist *Savante Arrhenius* in 1884.

Acids were defined as compounds that contain hydrogen and can dissolve in water to release hydrogen ions into solution. For example, HCl dissolves in water as follows:



Bases are defined as substances that dissolve in water to release hydroxide ions (OH^-) into solution. For example, a typical base according to the Arrhenius definition is sodium hydroxide, NaOH:



The process that converts a molecule such as HCl into ions is called ionization.

Ionization is the production of ions from molecular compounds. No ions were initially present.

The process of converting a solid ionic compound, such as NaOH, into ions in solution is called dissociation.

Dissociation is the production of ions from ionic compounds. Ions were initially present, water just breaks apart ions.

Some common acidic and basic substances:

Acidic	Basic
citrus fruits	soap
vinegar	detergents
aspirin	milk of magnesia

Some characteristic properties associated with acids and bases:

Acids	Bases
sour taste	bitter taste
change litmus paper to red	change litmus paper to blue
react with hydroxide bases to produce salt	slippery, soapy feeling

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