#### CHEMISTRY MODULE 2 – FOUNDATIONS IN CHEMISTRY

# 2.1 – ATOMS AND REACTIONS

#### Atomic structure and isotopes

 Describe the structure of an atom including the distribution of mass and charge.

	Protons	Electrons	Neutrons
Mass	1	1/2000th	1
Charge	+1	-1	0

- Mass is in nucleus(protons and neutons) with electron cloud surrounding it.
- Describe the difference between mass (nucleon) number and atomic (proton) number
  - Mass number the total number of protons and neutrons in the nucleus of an isotope
  - Atomic number the total number of protons in the nucleus of an atom
- Explain what an isotope is.
  - Atoms of an element (same number of protons) with different numbers of neutrons and different masses.
- Define the term 'first ionisation energy'
  - The amount of energy needed to remove an electron from each atom in 1 mole of gaseous atoms of an element.

#### Relative masses

- State what is used as the standard measurement for relative masses
  - o Carbon-12
- Define 'relative atomic mass' (Ar)
  - o The weighted average mass of an atom, relative to 1/12th the mass of a carbon-12 atom. THIS MEANS:

(The weighted average atomic mass of an element, relative to carbon 12, on a scale where carbon 12 is 12.)

- Define 'relative isotopic mass'
  - The mass of an isotope compared to 1/12th the mass of a carbon-12 atom.

THIS MEANS

(The mass of an isotope of an element relative to carbon 12 on a scale where carbon 12 is 12.)

# **Grade Boundaries Atoms Bonds Groups 2014**

Marcel A. Müller

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#### **Table of Contents Grade Boundaries Atoms Bonds Groups 2014**

- 1. Understanding the eBook Grade Boundaries Atoms Bonds Groups 2014
  - The Rise of Digital Reading Grade Boundaries Atoms Bonds Groups 2014
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Grade Boundaries Atoms Bonds Groups 2014
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Grade Boundaries Atoms Bonds Groups 2014
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Grade Boundaries Atoms Bonds Groups 2014
  - Personalized Recommendations
  - Grade Boundaries Atoms Bonds Groups 2014 User Reviews and Ratings
  - Grade Boundaries Atoms Bonds Groups 2014 and Bestseller Lists
- 5. Accessing Grade Boundaries Atoms Bonds Groups 2014 Free and Paid eBooks
  - Grade Boundaries Atoms Bonds Groups 2014 Public Domain eBooks
  - o Grade Boundaries Atoms Bonds Groups 2014 eBook Subscription Services
  - Grade Boundaries Atoms Bonds Groups 2014 Budget-Friendly Options
- 6. Navigating Grade Boundaries Atoms Bonds Groups 2014 eBook Formats
  - o ePub, PDF, MOBI, and More
  - o Grade Boundaries Atoms Bonds Groups 2014 Compatibility with Devices
  - Grade Boundaries Atoms Bonds Groups 2014 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Grade Boundaries Atoms Bonds Groups 2014
  - Highlighting and Note-Taking Grade Boundaries Atoms Bonds Groups 2014
  - Interactive Elements Grade Boundaries Atoms Bonds Groups 2014
- 8. Staying Engaged with Grade Boundaries Atoms Bonds Groups 2014

- Joining Online Reading Communities
- Participating in Virtual Book Clubs
- Following Authors and Publishers Grade Boundaries Atoms Bonds Groups 2014
- 9. Balancing eBooks and Physical Books Grade Boundaries Atoms Bonds Groups 2014
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Grade Boundaries Atoms Bonds Groups 2014
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Grade Boundaries Atoms Bonds Groups 2014
  - Setting Reading Goals Grade Boundaries Atoms Bonds Groups 2014
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Grade Boundaries Atoms Bonds Groups 2014
  - o Fact-Checking eBook Content of Grade Boundaries Atoms Bonds Groups 2014
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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