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**GRADE 11**

**MATHEMATICS**  
**PAPER 1 MARKING GUIDELINE**  
**JUNE EXAMINATION 2024**

# Maths 2014 Grade11 Exemplar Paper1

**Clifford Lane**



**Maths 2014 Grade11 Exemplar Paper1:**

## Unveiling the Power of Verbal Art: An Mental Sojourn through **Maths 2014 Grade11 Exemplar Paper1**

In some sort of inundated with displays and the cacophony of instant conversation, the profound energy and psychological resonance of verbal beauty usually diminish into obscurity, eclipsed by the regular onslaught of noise and distractions. Yet, nestled within the musical pages of **Maths 2014 Grade11 Exemplar Paper1**, a interesting function of literary splendor that pulses with organic emotions, lies an memorable trip waiting to be embarked upon. Published by a virtuoso wordsmith, this mesmerizing opus manuals viewers on a psychological odyssey, lightly exposing the latent potential and profound impact embedded within the complicated internet of language. Within the heart-wrenching expanse of this evocative examination, we can embark upon an introspective exploration of the book is central subjects, dissect its interesting publishing style, and immerse ourselves in the indelible impression it leaves upon the depths of readers souls.

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subsequent  $\Delta H$  values  $2 \text{ CO}_2(\text{g}) + \text{H}_2\text{O}(\text{g}) \rightarrow \text{C}_2\text{H}_2(\text{g}) + \frac{5}{2} \text{O}_2(\text{g})$   $\Delta H = 94.5 \text{ kJ}$   $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{O}_2(\text{g})$   $\Delta H = 71.2 \text{ kJ}$   $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2} \text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$   $\Delta H = 283 \text{ kJ}$   $\text{C}_2\text{H}_6(\text{g})$

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combustion of magnesium oxide  $\text{Mg} + \frac{1}{2}\text{O}_2 \rightarrow \text{MgO}$  we used

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = 55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}}$

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