



Quantitative Chemistry

Multiple Choice Questions

Set 1

Tick **one** box.

1. What is the correctly balanced symbol equation for the reaction between magnesium and hydrochloric acid?



2. What is the relative formula mass of dibromoethane ($\text{C}_2\text{H}_4\text{Br}_2$)?
Relative atomic masses (A_r): C = 12, H = 1, Br = 80

A. 93 ☐

B. 186 ☐

C. 188 ☐

D. 372 ☐

3. Which statement about moles is **not** true?

A. all substances contain the same number of moles ☐

B. one mole of carbon contains 6.02×10^{23} atoms of carbon ☐

C. the mass of one mole of a substance in grams is numerically equal to the relative formula mass of that substance ☐

D. the number of atoms, molecules or ions in one mole of a substance is the Avogadro constant ☐

4. What is the mass of two moles of rubidium?

Relative atomic mass (A_r) = 85

A. 42.5g ☐

B. 85g ☐

C. 127.5g ☐

D. 170g ☐

5. How many moles are there in 22g of carbon dioxide (CO_2)?

Relative atomic masses (A_r): C = 12, O = 16

A. 0.5 ☐

B. 2 ☐

C. 12 ☐

D. 968 ☐

Multiple Question For Chemistry

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Multiple Question For Chemistry:

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has been kept large enough to offer meaningful practice to the students **AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online Practice** Neil D. Jespersen, Pamela Kerrigan, 2023-07-04 A guide to taking the Advanced Placement exam in chemistry featuring a review of major chemistry concepts practice and diagnostic tests test taking strategies an overview of the test and practice problems **Study Guide to Accompany Basics for Chemistry** Martha Mackin, 2012-12-02 Study Guide to Accompany Basics for Chemistry is an 18 chapter text designed to be used with Basics for Chemistry textbook Each chapter contains Overview Topical Outline Skills and Common Mistakes which are all keyed to the textbook for easy cross reference The Overview section summarizes the content of the chapter and includes a comprehensive listing of terms a summary of general concepts and a list of numerical exercises while the Topical Outline provides the subtopic heads that carry the corresponding chapter and section numbers as they appear in the textbook The Fill in Multiple Choice are two sets of questions that include every concept and numerical exercise introduced in the chapter and the Skills section provides developed exercises to apply the new concepts in the chapter to particular examples The Common Mistakes section is designed to help avoid some of the errors that students make in their effort to learn chemistry while the Practical Test section includes matching and multiple choice questions that comprehensively cover almost every concept and numerical problem in the chapter After briefly dealing with an overview of chemistry this book goes on exploring the concept of matter energy measurement problem solving atom periodic table and chemical bonding These topics are followed by discussions on writing names and formulas of compounds chemical formulas and the mole chemical reactions calculations based on equations gases and the properties of a liquid The remaining chapters examine the solutions acids bases salts oxidation reduction reactions electrochemistry chemical kinetics and equilibrium and nuclear organic and biological chemistry This study guide will be of great value to chemistry teachers and students

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