

## The Hydrogen Atom – Student Challenge

### Background Material

Carefully read the background-information Energy Levels, Light, and Transitions, and answer the following questions as clearly (and completely) as you can.

Operation 1: Complete the following table which compares how the Bohr Model and the Quantum-mechanical approach the Hydrogen atom. In each “describing” box describe things as accurately as you can assuming they are true.

Concept	Bohr Model	Quantum Model
Electrons	positively charged particles in flat orbits	positively charged particles in flat orbits
Electrons	orbiting electrons	negatively charged clouds
Orbitals	Electrons do not move, they give the probability picture	Electrons, being particles without mass
Transitions of electrons in the atom orbital	moving a electron to other Bohr orbits	Electron wave
Energy of the electron levels	$E_1 = -13.6 \text{ eV}$ $E_2 = -3.4 \text{ eV}$	It is different

Operation 2: Complete the following table which compares the properties of two different elements by writing their appropriate relationships. The table now is completed for you: the red element has a higher frequency/energy, smaller energy, and the same number of protons as the other element.

# Hydrogen Atom Student Guide Solutions Naap

**Jessica J Manson**



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