## Gas Stoichiometry: Multiple Choice

 What volume of pure oxygen gas, O<sub>2</sub>, measured at 546 K and 101.325 kPa is formed by the complete decomposition of 0,200 mol Ag<sub>2</sub>O according to the following equation:

2Ag/O → 4Ag + O<sub>2</sub>

A) 1.12 L B) 4.48 L C) 2.24 L D) 8.96 L E) 17.82 L

ANS: B

I. How many moles of KClO<sub>2</sub> are needed to form 2.8 L of O<sub>2</sub>, measured at STP, according to the following reaction:  $2KClO_2 \rightarrow 2KCl + 3O_2$ 

A) 0.083 mol B) 0.25 mol C) 0.17 mol D) 0.50 mol E) 0.33 mol

ANS: A

1. What volume of chlorine gas, measured at 0°C and 101,325 kPa, is required to react completely with 5.61 g of KOH in the following reaction:  $Cl_{2m} + 2KOH \rightarrow KCI + KCIO + H_{2m}$ 

A) 0.56 L B) 4.48 L C) 1.12 L D) 22.4 L E) 2.24 L

ANS: C

Sodium hydride reacts with excess water to produce sodium hydroxide (aq) and hydrogen gas: NaH(s)
H (20) = NaCH(sex) + H (a)

+ H<sub>2</sub>O(l) → NaOH(aq) + H<sub>2</sub>(g)

A sample of Nall weighing \_\_\_\_\_ g will produce 982 mL of gas at 28.0 °C and 765 torr, when the hydrogen is collected over water. The vapor pressure of water at this temperature is 28 torr.

A) 2.93 B) 0.960 C) 0.925 D) 0.0388 E) 925

ANS: C

1. At STP, how many liters of oxygen are required to react completely with 3.6 liters of hydrogen to

form water?  $2H^{2}(g) + O^{2}(g) \rightarrow 2H^{2}O(g)$ 

A) 1.8 L B) 3.6 L C) 2.0 L D) 2.4 L

ANS: A

The molar volume of a gas at STP is \_\_\_\_\_\_ L.

A) 0.08206 B) 62.36 C) 1.00 D) 22.4 E) 14.7

ANS: D

 If 0.5 L of O<sub>2</sub>(g) reacts with H<sub>2</sub> to produce 1 L of H<sub>2</sub>O(g), what is the volume of H<sub>2</sub>O(g) obtained from 1 L of O<sub>3xi</sub>?

A) 0.5 L B) 1.5 L C) 2 L D) 2.5 L

ANS: C

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#### **Table of Contents Multiple Choice Questions On Chemical Stoichiometry**

- 1. Understanding the eBook Multiple Choice Questions On Chemical Stoichiometry
  - The Rise of Digital Reading Multiple Choice Questions On Chemical Stoichiometry
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Multiple Choice Questions On Chemical Stoichiometry
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Chemical Stoichiometry
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Multiple Choice Questions On Chemical Stoichiometry

- Personalized Recommendations
- Multiple Choice Questions On Chemical Stoichiometry User Reviews and Ratings
- Multiple Choice Questions On Chemical Stoichiometry and Bestseller Lists
- 5. Accessing Multiple Choice Questions On Chemical Stoichiometry Free and Paid eBooks
  - Multiple Choice Questions On Chemical Stoichiometry Public Domain eBooks
  - Multiple Choice Questions On Chemical Stoichiometry eBook Subscription Services
  - Multiple Choice Questions On Chemical Stoichiometry Budget-Friendly Options
- 6. Navigating Multiple Choice Questions On Chemical Stoichiometry eBook Formats
  - o ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Chemical Stoichiometry Compatibility with Devices
  - Multiple Choice Questions On Chemical Stoichiometry Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Chemical Stoichiometry
  - Highlighting and Note-Taking Multiple Choice Questions On Chemical Stoichiometry
  - Interactive Elements Multiple Choice Questions On Chemical Stoichiometry
- 8. Staying Engaged with Multiple Choice Questions On Chemical Stoichiometry
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Chemical Stoichiometry
- 9. Balancing eBooks and Physical Books Multiple Choice Questions On Chemical Stoichiometry
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Chemical Stoichiometry
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Multiple Choice Questions On Chemical Stoichiometry
  - Setting Reading Goals Multiple Choice Questions On Chemical Stoichiometry
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Multiple Choice Questions On Chemical Stoichiometry

- Fact-Checking eBook Content of Multiple Choice Questions On Chemical Stoichiometry
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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