

Three organic compounds contain carbon, hydrogen and oxygen only.

(a) An organic compound B has the following composition by mass:

$$\text{C, 66.7\%}, \text{H, 11.1\%}, \text{O, 22.2\%}$$

Calculate the empirical formula of compound B.

$$\text{C: } 66.7 \rightarrow \frac{66.7}{12} = 5.56 \rightarrow 5.56 \times 2 = 11.12 \rightarrow 11$$

$$\text{H: } 11.1 \rightarrow \frac{11.1}{1} = 11.1 \rightarrow 11.1 \times 2 = 22.2 \rightarrow 22$$

$$\text{O: } 22.2 \rightarrow \frac{22.2}{16} = 1.39 \rightarrow 1.39 \times 2 = 2.78 \rightarrow 3$$

$$\text{C: } 11 \rightarrow \text{C}_{11}\text{H}_{22}\text{O}_3$$

$$\text{H: } 22 \rightarrow \text{H}_{22}\text{O}_3$$

$$\text{O: } 3 \rightarrow \text{O}_3$$

$$\text{empirical formula} = \text{C}_{11}\text{H}_{22}\text{O}_3$$

(b) Compound B has the empirical formula $\text{C}_3\text{H}_6\text{O}$ and a relative molecular mass of 90.

Calculate the molecular formula of compound B.

$$\text{C}_3\text{H}_6\text{O} = 58$$

$$90 = 1.5 \times 58$$

$$\text{C} = 15$$

$$\frac{\text{relative mass}}{\text{empirical mass}} = \frac{90}{58} = 1.5 \quad \text{molecular formula} = \text{C}_5\text{H}_{10}\text{O}_2$$

$$\left[\begin{array}{l} 1.5 \times \text{C}_3\text{H}_6\text{O} = \text{C}_5\text{H}_{10}\text{O}_2 \\ \text{molecular formula} \end{array} \right.$$

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