

Date _____ Student Name Victor Carr
 Course/Section Chem 101-032 Team Members _____
 Instructor _____

90

The Identity of an Insoluble Precipitate

Pre-laboratory Assignment

1. The following reactions, shown in unbalanced equations, are pertinent to this experiment. Balance the equations:



2. When a solution containing 1.4 g of $\text{Ba}(\text{NO}_3)_2$ and 2.4 g of $\text{NH}_2\text{SO}_3\text{H}$ is boiled, a precipitate forms. This precipitate may be $\text{Ba}(\text{NH}_2\text{SO}_3)_2$, BaSO_4 , or $\text{Ba}(\text{NH}_2)_2$. You can find these products in the reactions whose equations you balanced in the first question.

- a. Calculate the number of moles of each reactant.

moles of Ba = $\frac{1.4 \text{ g}}{261.34 \text{ g/mol}} = 0.0054 \text{ moles}$?

moles of $\text{NH}_2\text{SO}_3\text{H}$ = $\frac{2.4 \text{ g}}{147.09 \text{ g/mol}} = 0.016 \text{ moles}$?

Identity Of An Insoluble Precipitate Prelab Answers

**Eugene LeMay, Jr., Herbert Beall, Karen
M. Robblee, Douglas C.
Brower, Prentice-Hall Staff**

Identity Of An Insoluble Precipitate Prelab Answers:

Chemistry Eugene LeMay, Jr.,Herbert Beall,Karen M. Robblee,Douglas C. Brower,Prentice-Hall Staff,2002-02

Chemistry in the Laboratory James M. Postma,Julian L. Robert,J. Leland Hollenberg,2004-03-12 This clearly written class tested manual has long given students hands on experience covering all the essential topics in general chemistry Stand alone experiments provide all the background introduction necessary to work with any general chemistry text This revised edition offers new experiments and expanded information on applications to real world situations **Exploring General**

Chemistry in the Laboratory Colleen F. Craig,Kim N. Gunnerson,2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **Chemical Investigations** Nancy Konigsberg Kerner,1986 **Methods in**

Biotechnology Seung-Beom Hong,M. Bazlur Rashid,Lory Z. Santiago-Vázquez,2016-05-12 As rapid advances in biotechnology occur there is a need for a pedagogical tool to aid current students and laboratory professionals in biotechnological methods Methods in Biotechnology is an invaluable resource for those students and professionals Methods in Biotechnology engages the reader by implementing an active learning approach provided advanced study questions as well as pre and post lab questions for each lab protocol These self directed study sections encourage the reader to not just perform experiments but to engage with the material on a higher level utilizing critical thinking and troubleshooting skills This text is broken into three sections based on level Methods in Biotechnology Advanced Methods in Biotechnology I and Advanced Methods in Biotechnology II Each section contains 14 22 lab exercises with instructor notes in appendices as well as an answer guide as a part of the book companion site This text will be an excellent resource for both students and laboratory professionals in the biotechnology field **Microscale and Miniscale Organic Chemistry Laboratory**

Experiments Allen M. Schoffstall,Barbara A. Gaddis,Melvin L. Druelinger,2000 This work offers a comprehensive introductory treatment of the organic laboratory techniques for handling glassware and equipment safety in the laboratory micro and mini scale experimental procedures theory of reactions and techniques applications and spectroscopy

Laboratory Experiments to Accompany General, Organic and Biological Chemistry David B. Macaulay,Joseph M. Bauer,Molly M. Bloomfield,2011 Organic chemists looking to build their understanding through lab work can utilize this second edition There are 21 experiments that are clearly described in the integrated table of contents Each one highlights the relevance and application of chemical principles to biological systems The experiments are designed to relate their

personal experience to the key concepts using common household and commercial products Each one is also written in an accessible way that assumes no prior work in the chemistry laboratory This makes it much easier for organic chemists to conduct each experiment and gain real world experience

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Table of Contents Identity Of An Insoluble Precipitate Prelab Answers

1. Understanding the eBook Identity Of An Insoluble Precipitate Prelab Answers
 - The Rise of Digital Reading Identity Of An Insoluble Precipitate Prelab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Identity Of An Insoluble Precipitate Prelab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Identity Of An Insoluble Precipitate Prelab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Identity Of An Insoluble Precipitate Prelab Answers
 - Personalized Recommendations
 - Identity Of An Insoluble Precipitate Prelab Answers User Reviews and Ratings

- Identity Of An Insoluble Precipitate Prelab Answers and Bestseller Lists
- 5. Accessing Identity Of An Insoluble Precipitate Prelab Answers Free and Paid eBooks
 - Identity Of An Insoluble Precipitate Prelab Answers Public Domain eBooks
 - Identity Of An Insoluble Precipitate Prelab Answers eBook Subscription Services
 - Identity Of An Insoluble Precipitate Prelab Answers Budget-Friendly Options
- 6. Navigating Identity Of An Insoluble Precipitate Prelab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Identity Of An Insoluble Precipitate Prelab Answers Compatibility with Devices
 - Identity Of An Insoluble Precipitate Prelab Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Identity Of An Insoluble Precipitate Prelab Answers
 - Highlighting and Note-Taking Identity Of An Insoluble Precipitate Prelab Answers
 - Interactive Elements Identity Of An Insoluble Precipitate Prelab Answers
- 8. Staying Engaged with Identity Of An Insoluble Precipitate Prelab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Identity Of An Insoluble Precipitate Prelab Answers
- 9. Balancing eBooks and Physical Books Identity Of An Insoluble Precipitate Prelab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Identity Of An Insoluble Precipitate Prelab Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Identity Of An Insoluble Precipitate Prelab Answers
 - Setting Reading Goals Identity Of An Insoluble Precipitate Prelab Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Identity Of An Insoluble Precipitate Prelab Answers
 - Fact-Checking eBook Content of Identity Of An Insoluble Precipitate Prelab Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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