

Qualitative Tests for Hydrocarbons, Alkanes, Alkenes, Alkynes

Hydrocarbons can be grouped into two structural classes and sample chemical tests can be used to distinguish between these classes. In this experiment you will perform four tests: combustion, bromine water, potassium permanganate, and Baeyer's test. Each using known hydrocarbons and their derivatives that were used as standards for structural analysis.

The four classes of hydrocarbons are as follows:

- (1) **Alkanes**—There are saturated hydrocarbons, usually single bonds.
- (2) **Alkenes**—There are hydrocarbons that contain at least one carbon-carbon double bond.
- (3) **Alkynes**—There are hydrocarbons that contain only isolated triple bonds.
- (4) **Aromatics**—There are hydrocarbons that contain no unsaturation. They possess rings in a special arrangement of alternating double and single bonds that are chemically stable due to the delocalization of electrons around the ring. Through the aromatic stabilization system is type stability towards other chemical reactions characteristic of the representative. They are small molecules that possess double bonds. They are due to the fact that they are not easily broken down by a reaction of double bond rings. Aromatic is a very specific term. There are four types of the following: benzene, toluene, xylene, and naphthalene.



Put the test tubes into an holder, the hydrocarbon holder as Figure 11.2.



Test the hydrocarbons with following tests

The combustion of alkanes and alkenes can be easily tested with a flame test and other methods of test. The combustion of the alkane is easily distinguished by the flame test. The combustion of the alkene is easily distinguished by the flame test. The combustion of the alkane is easily distinguished by the flame test. The combustion of the alkene is easily distinguished by the flame test.

Procedure:

Sample 1: alkane, cyclohexane. Add about 1 mL of cyclohexane to a test tube. Add a few drops of bromine. The solution is colorless. Add a few drops of potassium permanganate. The solution is colorless. Add a few drops of Baeyer's reagent. The solution is colorless. Sample 2: alkene, cyclohexene. Add about 1 mL of cyclohexene to a test tube. Add a few drops of bromine. The solution is colorless. Add a few drops of potassium permanganate. The solution is colorless. Add a few drops of Baeyer's reagent. The solution is colorless. Sample 3: alkyne, cyclohexadiene. Add about 1 mL of cyclohexadiene to a test tube. Add a few drops of bromine. The solution is colorless. Add a few drops of potassium permanganate. The solution is colorless. Add a few drops of Baeyer's reagent. The solution is colorless.



Hydrocarbon Lab Manual

RC Schank

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

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