

Arrhenius' concept based on water

Arrhenius, 1880s:

Acids form hydrogen ions $H^+(H_2O)_n$ in aqueous solution.

Bases form hydroxide ions in aqueous solution.

Examples of Arrhenius acids (in water): HCl , H_2SO_4 , etc.

Examples of Arrhenius bases (in water): $NaOH$, NH_3 , etc.

Arrhenius definitions only apply to aqueous solutions.

A general Arrhenius acid-base reaction is the reaction between H^+ and OH^- to produce water.

A Neutralization Reaction



Guided Reading And Study Workbook Chapter 20 Acids And Bases

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