

## Section 2–1 The Nature of Matter (pages 35–39)

This section identifies the three particles that make up atoms. It also explains how atoms of the same element can have a different number of neutrons and describes the two main types of chemical bonds.

### Atoms (page 35)

1. The basic unit of matter is called a(an) .
2. Describe the nucleus of an atom.
3. Complete the table about subatomic particles.

SUBATOMIC PARTICLES		
Particle	Charge	Location in Atom
	Positive	
	Neutral	
	Negative	

4. Why are atoms neutral despite having charged particles?

### Elements (page 36)

5. What is a chemical element?
6. What does an element's atomic number represent?

### Chemical Compounds (page 37)

7. What is a chemical compound?
8. What are the elements in salt? How many of each element are there?

### Chemical Bonds (pages 38–39)

9. What holds atoms in compounds together?
10. Complete the table about the main types of chemical bonds.

## CHEMICAL BONDS

Chemical bonds	
Type	Formed When...
Covalent bond	
Ionic bond	

11. What is an ion?
12. Is the following sentence **true** or **false**?  
An atom that loses electrons has a negative charge.
13. The structure that results when atoms are joined together by covalent bonds is called a(an) .

# Guided Reading And Study Workbook Chapter 2answers

**L Manion**



## **Guided Reading And Study Workbook Chapter 2answers:**

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## **Table of Contents Guided Reading And Study Workbook Chapter 2answers**

1. Understanding the eBook Guided Reading And Study Workbook Chapter 2answers
  - The Rise of Digital Reading Guided Reading And Study Workbook Chapter 2answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Guided Reading And Study Workbook Chapter 2answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Guided Reading And Study Workbook Chapter 2answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Guided Reading And Study Workbook Chapter 2answers
  - Personalized Recommendations
  - Guided Reading And Study Workbook Chapter 2answers User Reviews and Ratings

- Guided Reading And Study Workbook Chapter 2answers and Bestseller Lists
- 5. Accessing Guided Reading And Study Workbook Chapter 2answers Free and Paid eBooks
  - Guided Reading And Study Workbook Chapter 2answers Public Domain eBooks
  - Guided Reading And Study Workbook Chapter 2answers eBook Subscription Services
  - Guided Reading And Study Workbook Chapter 2answers Budget-Friendly Options
- 6. Navigating Guided Reading And Study Workbook Chapter 2answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Guided Reading And Study Workbook Chapter 2answers Compatibility with Devices
  - Guided Reading And Study Workbook Chapter 2answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Guided Reading And Study Workbook Chapter 2answers
  - Highlighting and Note-Taking Guided Reading And Study Workbook Chapter 2answers
  - Interactive Elements Guided Reading And Study Workbook Chapter 2answers
- 8. Staying Engaged with Guided Reading And Study Workbook Chapter 2answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Guided Reading And Study Workbook Chapter 2answers
- 9. Balancing eBooks and Physical Books Guided Reading And Study Workbook Chapter 2answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Guided Reading And Study Workbook Chapter 2answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Guided Reading And Study Workbook Chapter 2answers
  - Setting Reading Goals Guided Reading And Study Workbook Chapter 2answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Guided Reading And Study Workbook Chapter 2answers
  - Fact-Checking eBook Content of Guided Reading And Study Workbook Chapter 2answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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