

Unit 4: Stoichiometry

- d. How many moles of NH_3 will be produced if 0.2 moles of CaCN_2 are used in the reaction?

$$0.2 \text{ moles } \text{CaCN}_2 \times \frac{2 \text{ mol } \text{NH}_3}{1 \text{ mol } \text{CaCN}_2} = 1 \text{ mol } \text{NH}_3$$



- a. What is the mole ratio of Cl_2 to PCl_5 ?

$$\frac{1 \text{ mol } \text{Cl}_2}{1 \text{ mol } \text{PCl}_5}$$

- b. How many moles of Cl_2 will react with 0.404 moles of PCl_5 ?

$$0.404 \text{ mol } \text{PCl}_5 \times \frac{1 \text{ mol } \text{Cl}_2}{1 \text{ mol } \text{PCl}_5} = 0.404 \text{ mol } \text{Cl}_2$$

Gram-to-Gram Stoichiometry



- a. What is the molar mass of O_2 ?

$$2(16.00) = 32.00 \text{ g}$$

- b. What is the molar mass of H_2O ?

$$2(18.02) = 36.04 \text{ g}$$

- c. What is the mole-to-mole ratio of O_2 to H_2O ?

$$\frac{5 \text{ mol } \text{O}_2}{2 \text{ mol } \text{H}_2\text{O}}$$

- d. How many grams of O_2 reacted, if 17.4 grams of H_2O are produced?

$$17.4 \text{ g } \text{H}_2\text{O} \times \frac{1 \text{ mol } \text{H}_2\text{O}}{18.02 \text{ g } \text{H}_2\text{O}} \times \frac{5 \text{ mol } \text{O}_2}{2 \text{ mol } \text{H}_2\text{O}} = 24.08 \text{ g } \text{O}_2$$

$$17.4 \text{ g } \text{H}_2\text{O} \times \frac{1 \text{ mol } \text{H}_2\text{O}}{18.02 \text{ g } \text{H}_2\text{O}} \times \frac{5 \text{ mol } \text{O}_2}{2 \text{ mol } \text{H}_2\text{O}} \times \frac{32.00 \text{ g } \text{O}_2}{1 \text{ mol } \text{O}_2} = 24.08 \text{ g } \text{O}_2$$

Guided Practice Problems 1stoichiometry

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R.N.,Robin Kavanagh**



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