## CHEMISTRY 10TH - Multiple Choice Questions



CHAPTER# 01

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- 01. Which one of the following statements is false about dynamic equilibrium?
  - (a) It takes place in a closed container
  - (b) Concentration of reactant and products are not changed
  - (c) Rate of the forward reaction is equal to the rate of the reverse reaction
  - (d) ▼ Equilibrium cannot be disturbed by any external stress
- 02. Consider the following reaction and indicate which of the following best describes equilibrium constant expression  $K_c$ .  $4NH_3 + 5O_2 = 4NO + 6H_2O$

(a) 
$$V K_c = \frac{[NO]^4 [H_2O]^6}{[NH_3]^4 [O_2]^5}$$

(c) 
$$K_c = \frac{[NH_3][O_2]}{[H_1O][NO]}$$

(b) 
$$K_c = \frac{[NH_3]^4 [O_2]^5}{[NO]^4 [H_2O]^6}$$

(d) 
$$K_c = \frac{[4NO][6H_2O]}{[4NH_3][5O_2]}$$

- O3. A reaction which is never goes to completion is known as reversible reaction. Reversible reaction is represented by:
  - (a) doted lines

(b) single arrow

(c) V double arrow

- (d) double straight line
- 04. When the magnitude of K<sub>c</sub> is small, indicates:
  - (a) V Reaction mixture contain most of the reactant
  - (b) Reaction mixture contain most of the product
  - (c) Reaction mixture contain almost equal amount of reactant and product
  - (d) Reaction goes to completion
- 05. For which system does the equilibrium constant, K<sub>c</sub> has units of concentration?

(b) 
$$N_{2(g)} + O_{2(g)} \Leftrightarrow 2NO_{(g)}$$

(c) 
$$H_{2(g)} + I_{2(g)} \Leftrightarrow 2HI_{(g)}$$

(d) 
$$CO_{2(g)} + H_{2(g)} \Rightarrow CO_{(g)} + H_2O_{(1)}$$

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**R Bogdan** 

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