

IDEAL GAS LAW

Name: _____

Use the Ideal Gas Law below to solve the following problems:

$$PV = nRT \quad \text{where } P = \text{pressure in atmospheres}$$

$$V = \text{volume in liters}$$

$$n = \text{number of moles of gas}$$

$$R = \text{Universal Gas Constant}$$

$$0.0821 \text{ L} \cdot \text{atm} / \text{mol} \cdot \text{K}$$

$$T = \text{Kelvin temperature}$$

1. A gas sample with a volume of 2.50 L at 1.00 atm and 25°C is compressed to 1.25 L at 2.00 atm and 25°C.

2. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C.

3. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

4. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

5. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

6. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

7. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

8. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

9. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

10. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

11. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

12. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

13. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is compressed to 0.50 L at 2.00 atm and 25°C. The gas is nitrogen.

14. A gas sample with a volume of 1.00 L at 1.00 atm and 25°C is expanded to 2.00 L at 0.50 atm and 25°C. The gas is nitrogen.

Ideal Gas Law Packet Answer Key

VM Jensen



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