

19.2

HYDROGEN IONS AND ACIDITY

Section Review

Objectives

1. Identify a conjugate acid-base pair in chemical reactions involving acids and bases.
2. Calculate the pH of a solution from the concentration of H^+ or H_3O^+ ions, and vice versa.
3. Determine the pH of a solution from K_a and K_b .

Vocabulary

acid	1. conjugate base	acid solution
acid-base indicator	2. hydronium ion	acidic solution
acidic solution	3. autoionization	pH

Key Equation

- a. $pH = -\log [H^+]$ (ppm 1.4.1 and 1.4.2)
- b. $[H^+] = 10^{-pH}$ (ppm 1.4.1 and 1.4.2)

Part A Completion

Write the answers to the following questions in the spaces provided. Show necessary calculations and units. Write answers in complete sentences.

1. Acetic acid, CH_3COOH , is a weak acid. Write the chemical equation for its autoionization. (ppm 1.4.1)
2. Write the chemical equation for the autoionization of water. (ppm 1.4.1)
3. The pH of a solution is 4.5. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
4. The pH of a solution is 9.2. What is the concentration of OH^- ions in this solution? (ppm 1.4.1)
5. The pH of a solution is 3.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
6. The pH of a solution is 11.0. What is the concentration of OH^- ions in this solution? (ppm 1.4.1)
7. The pH of a solution is 2.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
8. The pH of a solution is 12.0. What is the concentration of OH^- ions in this solution? (ppm 1.4.1)
9. The pH of a solution is 1.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
10. The pH of a solution is 13.0. What is the concentration of OH^- ions in this solution? (ppm 1.4.1)
11. The pH of a solution is 0.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
12. The pH of a solution is 14.0. What is the concentration of OH^- ions in this solution? (ppm 1.4.1)
13. The pH of a solution is 5.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
14. The pH of a solution is 6.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
15. The pH of a solution is 7.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
16. The pH of a solution is 8.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
17. The pH of a solution is 9.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
18. The pH of a solution is 10.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
19. The pH of a solution is 11.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
20. The pH of a solution is 12.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
21. The pH of a solution is 13.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)
22. The pH of a solution is 14.0. What is the concentration of H^+ ions in this solution? (ppm 1.4.1)

Hydrogen Ions And Acidity Section Review Answers

AW Rasmussen



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