

Name: _____ Period: _____

Temperature Conversions and Heat Practice

Please show all work to receive credit

$$K = ^\circ C + 273$$

$$^\circ F = (9/5 \times ^\circ C) + 32$$

$$^\circ C = 5/9(^{\circ}F - 32)$$

Convert the following to Fahrenheit

1) $10^{\circ}C$ _____

2) $30^{\circ}C$ _____

3) $0^{\circ}C$ _____

Convert the following to Celsius

4) $32^{\circ}F$ _____

5) $90^{\circ}F$ _____

6) $212^{\circ}F$ _____

Convert the following to Kelvin

7) $0^{\circ}C$ _____

8) $-50^{\circ}C$ _____

9) $90^{\circ}C$ _____

10) $-20^{\circ}C$ _____

Convert the following to Celsius

11) $112^{\circ}K$ _____

12) $200^{\circ}K$ _____

13) $273^{\circ}K$ _____

14) $350^{\circ}K$ _____

Use the information below to complete the following problems

Heat = mass x specific heat x change in temperature

Δt = final temperature – initial temperature

$$Q = m c \Delta t$$

$$\Delta t = t_f - t_i$$

Substances	(j/kg•K)	Substance	(j/kg•K)	Substance	(j/kg•K)
Water (liquid)	4186	Copper	385	Iron	449
Steam	1870	Gold	129	Aluminum	897

15) How much energy is needed to heat 55 Kg of water from 285 K to 360 K?

16) How much energy is released when 55 Kg of steam cools from 558 K to 390 K?

17) If 5785 J of energy is used to heat 48 kg of Iron what is it's change in temperature?

18) An 75 kg block of metal absorbs 721875 J of heat as the temperature increases from 375 K to 400 K.
What is the specific heat of the metal?

a. Which metal is it likely to be?

19) How much energy is released if 278 kG of Aluminum cools from 698 K to 298 K?

20) What is the specific heat of a 3.78 kg object that absorbs 678 J as the temperature increases by 4.25 K?

Het And Its Measurement Answer Key

Yan Bai



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Table of Contents Het And Its Measurement Answer Key

1. Understanding the eBook Het And Its Measurement Answer Key
 - The Rise of Digital Reading Het And Its Measurement Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Het And Its Measurement Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Het And Its Measurement Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Het And Its Measurement Answer Key
 - Personalized Recommendations

- Het And Its Measurement Answer Key User Reviews and Ratings
- Het And Its Measurement Answer Key and Bestseller Lists
- 5. Accessing Het And Its Measurement Answer Key Free and Paid eBooks
 - Het And Its Measurement Answer Key Public Domain eBooks
 - Het And Its Measurement Answer Key eBook Subscription Services
 - Het And Its Measurement Answer Key Budget-Friendly Options
- 6. Navigating Het And Its Measurement Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Het And Its Measurement Answer Key Compatibility with Devices
 - Het And Its Measurement Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Het And Its Measurement Answer Key
 - Highlighting and Note-Taking Het And Its Measurement Answer Key
 - Interactive Elements Het And Its Measurement Answer Key
- 8. Staying Engaged with Het And Its Measurement Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Het And Its Measurement Answer Key
- 9. Balancing eBooks and Physical Books Het And Its Measurement Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Het And Its Measurement Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Het And Its Measurement Answer Key
 - Setting Reading Goals Het And Its Measurement Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Het And Its Measurement Answer Key
 - Fact-Checking eBook Content of Het And Its Measurement Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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