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Chapter 8 Atomic Structure Section 8.1 Studying Atomic (pages 189-195) This section explains how atomic models evolved. **Reading Strategy (page 190)** Summarizing As you read, fill in the table about atomic models. For more details on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of your textbook. **Atomic Models Scientific Evidence Model** Models of atoms in compounds Deducted from Rutherford's Positive, dense nucleus. **Atomic Greek Models of Atoms (page 191)** 1. Democritus called the smallest particles of matter because they could not be split. 2. He saw the four elements that Aristotle used in his model of matter: a, b, c, d. Dalton's Atomic Theory (page 192) 3. Thomson's Model

Model	Structure
1. Democritus's Model	Democritus proposed that matter is made of very small particles called atoms. Atoms cannot be created or destroyed. Atoms are indivisible particles, which cannot be created or destroyed.
2. Dalton's Atomic Theory	Dalton's Atomic Theory is based on the following postulates: (i) Matter is made of very small particles called atoms. (ii) Atoms cannot be created or destroyed. (iii) Atoms are indivisible particles, which cannot be created or destroyed.
3. Thomson's Model	Thomson's model of an atom is a sphere of uniform positive charge with electrons embedded in it. The electrons are negatively charged particles. The overall charge of the atom is zero because the positive and negative charges are equal in magnitude.

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