



COVALENT BONDING

Practice Problems

Do your answers to each of the following problems.

SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride b. Nitrogen dioxide c. Hydrogen peroxide
d. Carbon disulfide e. Sulfur dioxide

2. Write the formula for the following molecular compounds.

- a. Carbon monoxide b. Dinitrogen pentoxide c. Hydrogen cyanide
d. Nitrogen dioxide e. Sulfur trioxide

3. Draw a Lewis structure for each of the following molecular compounds.

a. Dichloromethane b. Carbon tetrachloride c. Hydrogen peroxide d. Hydrogen cyanide e. Hydrogen sulfide f. Hydrogen chloride

4. Draw a Lewis structure for each of the following molecular compounds. Label each atom as being hypervalent, electron deficient, or obeying the octet rule.

SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw a Lewis structure for hydrogen fluoride, HF.



2. Draw a Lewis structure for hydrogen peroxide, H_2O_2 .



3. Draw a Lewis structure for carbon dioxide, CO_2 .



4. Draw a Lewis structure for acetylene, C_2H_2 .

5. Draw a Lewis structure for each of the following. Label each atom as being hypervalent, electron deficient, or obeying the octet rule.



Obeying octet:

SECTION 8.3 BONDING IN THE LARSEN

1. Write the Lewis structure for each of the following compounds. Label each atom as being hypervalent, electron deficient, or obeying the octet rule.

2. Write the Lewis structure for each of the following. Label each atom as being hypervalent, electron deficient, or obeying the octet rule.

Guided Practice Problem Chapter 8 Covalent Bonding

Arshad Iqbal



Guided Practice Problem Chapter 8 Covalent Bonding:

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